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Date: **October 4, 2012**

Student Exploration: Electron Configuration

Activity B: Atomic radii

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s2 2s2 2p6 3s1**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s2 2s2 2p6 3s2**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s2 2s2 2p6 3s2 3p1	118
Silicon	14	1s2 2s2 2p6 3s2 3p2	111
Phosphorus	15	1s2 2s2 2p6 3s2 3p3	98
Sulfur	16	1s2 2s2 2p6 3s2 3p4	88
Chlorine	17	1s2 2s2 2p6 3s2 3p5	79
Argon	18	1s2 2s2 2p6 3s2 3p6	71

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Table of Contents Explore Learning Student Exploration Electron Configuration Answers

1. Understanding the eBook Explore Learning Student Exploration Electron Configuration Answers
 - The Rise of Digital Reading Explore Learning Student Exploration Electron Configuration Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Student Exploration Electron Configuration Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Student Exploration Electron Configuration Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Student Exploration Electron Configuration Answers
 - Personalized Recommendations
 - Explore Learning Student Exploration Electron Configuration Answers User Reviews and Ratings
 - Explore Learning Student Exploration Electron Configuration Answers and Bestseller Lists
5. Accessing Explore Learning Student Exploration Electron Configuration Answers Free and Paid eBooks
 - Explore Learning Student Exploration Electron Configuration Answers Public Domain eBooks
 - Explore Learning Student Exploration Electron Configuration Answers eBook Subscription Services
 - Explore Learning Student Exploration Electron Configuration Answers Budget-Friendly Options

6. Navigating Explore Learning Student Exploration Electron Configuration Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Student Exploration Electron Configuration Answers Compatibility with Devices
 - Explore Learning Student Exploration Electron Configuration Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Student Exploration Electron Configuration Answers
 - Highlighting and Note-Taking Explore Learning Student Exploration Electron Configuration Answers
 - Interactive Elements Explore Learning Student Exploration Electron Configuration Answers
8. Staying Engaged with Explore Learning Student Exploration Electron Configuration Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Student Exploration Electron Configuration Answers
9. Balancing eBooks and Physical Books Explore Learning Student Exploration Electron Configuration Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Student Exploration Electron Configuration Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Student Exploration Electron Configuration Answers
 - Setting Reading Goals Explore Learning Student Exploration Electron Configuration Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Student Exploration Electron Configuration Answers
 - Fact-Checking eBook Content of Explore Learning Student Exploration Electron Configuration Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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