

Moorest College
Chemistry 11 Mock Final Examination **ANSWERS**
Professor Torres

1. What is the pH of a solution that is $1.7 \times 10^{-2} \text{ M H}^+$?
A. **4.77**
B. 10.98
C. 0.22
D. 9.23
E. 1.70
2. Which combination below will form a buffer solution?
A. HCl and NaCl
B. HNO_3 and NaNO_3
C. **$\text{HC}_2\text{H}_3\text{O}_2$ and $\text{NaC}_2\text{H}_3\text{O}_2$**
D. NaBr and NaOH
E. All of the above are true.
3. A sample of helium gas initially at 37.0°C , 785 torr, and 2.00 L was heated to 58.0°C while the volume expanded to 3.24 L. What is the final pressure in atm?
A. 517
B. **0.681**
C. 1.79
D. 3.21
E. None of the above
4. Consider the reaction of an aqueous solution of lead(II) nitrate with an aqueous solution of potassium iodide. Which of the following represents the correctly balanced net ionic equation?
A. $\text{Pb}^{2+}(\text{aq}) + \text{I}_2(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$
B. $2\text{NO}_3^-(\text{aq}) + 2\text{K}^+(\text{aq}) \rightarrow 2\text{KNO}_3$
C. $\text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) + 2\text{K}^+(\text{aq}) + 2\text{I}^-(\text{aq}) \rightarrow \text{PbI}_2(\text{s}) + 2\text{K}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$
D. **$\text{Pb}^{2+}(\text{aq}) + 2\text{I}^-(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$**
E. None of the above
5. What is the molecular geometry of sulfur dioxide, SO_2 ?
A. **bent**
B. linear
C. tetrahedral
D. trigonal pyramidal
E. octahedral

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Table of Contents Moorpark Chemistry Answer Sheets

1. Understanding the eBook Moorpark Chemistry Answer Sheets
 - The Rise of Digital Reading Moorpark Chemistry Answer Sheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Moorpark Chemistry Answer Sheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Moorpark Chemistry Answer Sheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Moorpark Chemistry Answer Sheets
 - Personalized Recommendations
 - Moorpark Chemistry Answer Sheets User Reviews and Ratings
 - Moorpark Chemistry Answer Sheets and Bestseller Lists
5. Accessing Moorpark Chemistry Answer Sheets Free and Paid eBooks
 - Moorpark Chemistry Answer Sheets Public Domain eBooks
 - Moorpark Chemistry Answer Sheets eBook Subscription Services
 - Moorpark Chemistry Answer Sheets Budget-Friendly Options

6. Navigating Moorpark Chemistry Answer Sheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Moorpark Chemistry Answer Sheets Compatibility with Devices
 - Moorpark Chemistry Answer Sheets Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Moorpark Chemistry Answer Sheets
 - Highlighting and Note-Taking Moorpark Chemistry Answer Sheets
 - Interactive Elements Moorpark Chemistry Answer Sheets
8. Staying Engaged with Moorpark Chemistry Answer Sheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Moorpark Chemistry Answer Sheets
9. Balancing eBooks and Physical Books Moorpark Chemistry Answer Sheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Moorpark Chemistry Answer Sheets
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Moorpark Chemistry Answer Sheets
 - Setting Reading Goals Moorpark Chemistry Answer Sheets
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Moorpark Chemistry Answer Sheets
 - Fact-Checking eBook Content of Moorpark Chemistry Answer Sheets
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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