

$$\sin x = 0$$

$$\Rightarrow x = n\pi, n \in \mathbb{I}$$

$$\sin 2x = 0$$

$$\Rightarrow 2x = n\pi$$

$$\Rightarrow x = (n\pi/2), n \in \mathbb{I}$$

$$\sin 3x = 0$$

$$\Rightarrow 3x = n\pi$$

$$\Rightarrow x = (n\pi/3), n \in \mathbb{I}$$

$x = n\pi/2$  does not satisfy the given equation.

Hence, the required solution is  $x = n\pi$  and  $x = n\pi/3$ .

**OR**

In a triangle, ABC, prove that  $\sin A + \sin B + \sin C = 4 \cos A/2 \cos B/2 \cos C/2$

**Solution:**

In any triangle ABC,

$$A + B + C = \pi$$

$$A + B = \pi - C$$

$$(A + B)/2 = (\pi - C)/2$$

$$(A + B)/2 = (\pi/2) - (C/2)$$

$$\text{LHS} = \sin A + \sin B + \sin C$$

$$= 2 \sin [(A + B)/2] \cos [(A - B)/2] + \sin C$$

$$= 2 \sin [(\pi/2) - (C/2)] \cos [(A - B)/2] + 2 \sin (C/2) \cos (C/2)$$

$$= 2 \cos (C/2) \cos [(A - B)/2] + 2 \sin (C/2) \cos (C/2)$$

$$= 2 \cos (C/2) [\cos [(A - B)/2] + \sin (C/2)]$$

$$= 2 \cos \left( \frac{C}{2} \right) \left\{ \cos \left( \frac{A - B}{2} \right) + \sin \left( \frac{\pi - (A + B)}{2} \right) \right\}$$

$$= 2 \cos \left( \frac{C}{2} \right) \left\{ \cos \left( \frac{A - B}{2} \right) + \sin \left( \frac{\pi}{2} - \frac{A + B}{2} \right) \right\}$$

$$= 2 \cos \left( \frac{C}{2} \right) \left\{ \cos \left( \frac{A - B}{2} \right) + \cos \left( \frac{A + B}{2} \right) \right\}$$

$$= 2 \cos \left( \frac{C}{2} \right) \left\{ 2 \cos \left( \frac{\frac{A - B}{2} + \frac{A + B}{2}}{2} \right) \cos \left( \frac{\frac{A - B}{2} - \frac{A + B}{2}}{2} \right) \right\}$$

$$= 2 \cos (C/2) [2 \cos (A/2) \cos (-B/2)]$$

$$= 4 \cos (A/2) \cos (B/2) \cos (C/2)$$

$$= \text{RHS}$$

Hence proved.

#### Question 4

Prove that:  $5^{2n} - 1$  is divisible by 24 for all  $n \in \mathbb{N}$ .

[4]

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

web question experiment 8 determining ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 hci molarity 0.050 m volume of filtered ca oh 2 solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml oh in saturated ca oh 2 m ca in saturated ca oh 2 m ksp of ca oh 2 average ksp of ca oh 2

the chemical equation used was  $\text{Ca(OH)}_2 \text{ (s)} + 2\text{HCl (aq)} \rightarrow \text{CaCl}_2 \text{ (aq)} + 2\text{H}_2\text{O (l)}$  the solubility product was determined to be  $K_{\text{sp}} \text{ Ca(OH)}_2$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3(\text{PO}_4)_2$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore

$$K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$$

web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated  $\text{Ca(OH)}_2$  soultion with a st andard hcl solution and find the  $K_{sp}$  or the  $\text{OH}^-$  concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

web question determining the ksp of calciumhydroxide  $\text{Ca(OH)}_2$  s  $\text{Ca}^{2+}$  aq  $2\text{OH}^-$  aq for this experiment the titration of calciumhydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide  $\text{Ca(OH)}_2$  s  $\text{Ca}^{2+}$  aq  $2\text{OH}^-$  aq for this experiment the titration of calcium

web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of  $\text{Ca(OH)}_2$  is represented in equation form

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