CHEMISTRY PAPER 1 (THEORY)

Maximum Marks: 70

Time Allowed: Three hours

Candidates are allowed additional 15 minutes for only reading the paper.

They must NOT start writing during this time.

This paper is divided into four sections - A. B. C and D.

Answer all questions.

Section A consists of one question having sub-parts of one mark each.

Section B consists of ten questions of two marks each.

Section C consists of seven questions of three marks each, and

Section D consists of three questions of five marks each.

Internal choices have been provided in one question each in Section B, Section C and Section D.

All working, including rough work, should be done on the same sheet as, and adjacent to the rest of the answer.

The intended marks for questions or parts of questions are given in brackets []. Balanced equations must be given wherever possible and diagrams where they are helpful.

When solving numerical problems, all essential working must be shown.

In working out problems, use the following data:

Gas constant R = 1.987 cal deg' mot' = 8.314 JK' mot' = 0.0821 dm' atm K'mot'

I I atm = I dm3 atm = 1013 J. I Faraday = 96500 coulombs.

Avogadro's number = 6.023×10^{19} .

SECTION A – 14 MARKS

Ouestion 1

(A) Fill in the blanks by choosing the appropriate word(s) from those given in the [4×1] brackets:

[stable, low, aldehyde, unstable, 6, 4, ethane, Clemmensen's, 2, 3, carboxylic acid, high, propane, Rosenmund's]

- (i) The primary alcohols are easily oxidised first into _____ and then into
- (ii) The intermediate activated complex in a chemical reaction is highly _____ due to _____ energy.
- (iii) The coordination number and exidation state of the complex K₄[Fe(CN)₆] are _____ and ____ respectively.
- (iv) Propanone on reaction with zinc-amalgam in presence of conc. HCl gives and the reaction is known as _____ reduction.

This Paper consists of 8 printed pages.

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Grade Chemistry Third Term Exam Paper

Michael Brown

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