

CHAPTER 21 REVIEW

Nuclear Chemistry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ Based on the information about the three elementary particles in the text, which has the greatest mass?
 - (a) the proton
 - (b) the neutron
 - (c) the electron
 - (d) They all have the same mass.
2. _____ The force that keeps nucleons together is
 - (a) a strong nuclear force.
 - (b) a weak nuclear force.
 - (c) an electromagnetic force.
 - (d) a gravitational force.
3. _____ The stability of a nucleus is most affected by the
 - (a) number of neutrons.
 - (b) number of protons.
 - (c) number of electrons.
 - (d) ratio of neutrons to protons.
4. _____ If an atom should form from its constituent particles,
 - (a) matter is lost and energy is taken in.
 - (b) matter is lost and energy is released.
 - (c) matter is gained and energy is taken in.
 - (d) matter is gained and energy is released.
5. _____ For atoms of a given mass number, those with greater mass defects, have
 - (a) smaller binding energies per nucleon.
 - (b) greater binding energies per nucleon.
 - (c) the same binding energies per nucleon as those with smaller mass defects.
 - (d) variable binding energies per nucleon.
6. Based on **Figure 1.1** of the text, which isotope of He, helium-3 or helium-4,
_____ a. has the smaller binding energy per nucleon?
_____ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is _____.
8. Nuclides of the same element have the same _____.

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web mesoscopic physics is a fast growing discipline with countless potential applications understanding the science of mesoscopic materials with unique physical and chemical properties is important for the design of nanodevices and materials with unique properties

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web mesoscopic physics is a subdiscipline of condensed matter physics that deals with materials of an intermediate size these materials range in size between the nanoscale for a quantity of atoms such as a molecule and of materials measuring micrometres 1 the lower limit can also be defined as being the size of individual atoms

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