



55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.

Solution:

1. White phosphorus

(i) It is soft waxy solid having garlic odour

(ii) It is poisonous

(iii) Has low melting point and boiling point because P_4 molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.

Solution:



57. PCl_5 reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH_3 solution. Write the reactions involved to explain what happens.

Solution:

PCl_5 reacts with silver to form white silver salt ($AgCl$). Which then dissolves on adding excess aqueous NH_3 to form a soluble complex



58. Phosphorus forms a number of oxoacids. Out of these oxoacids phosphinic acid has strong

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Table of Contents Glencoe Chemistry Text Answers

1. Understanding the eBook Glencoe Chemistry Text Answers
 - The Rise of Digital Reading Glencoe Chemistry Text Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Text Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glencoe Chemistry Text Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Text Answers

- Personalized Recommendations
- Glencoe Chemistry Text Answers User Reviews and Ratings
- Glencoe Chemistry Text Answers and Bestseller Lists
- 5. Accessing Glencoe Chemistry Text Answers Free and Paid eBooks
 - Glencoe Chemistry Text Answers Public Domain eBooks
 - Glencoe Chemistry Text Answers eBook Subscription Services
 - Glencoe Chemistry Text Answers Budget-Friendly Options
- 6. Navigating Glencoe Chemistry Text Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Glencoe Chemistry Text Answers Compatibility with Devices
 - Glencoe Chemistry Text Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glencoe Chemistry Text Answers
 - Highlighting and Note-Taking Glencoe Chemistry Text Answers
 - Interactive Elements Glencoe Chemistry Text Answers
- 8. Staying Engaged with Glencoe Chemistry Text Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glencoe Chemistry Text Answers
- 9. Balancing eBooks and Physical Books Glencoe Chemistry Text Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glencoe Chemistry Text Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Glencoe Chemistry Text Answers
 - Setting Reading Goals Glencoe Chemistry Text Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Glencoe Chemistry Text Answers

- Fact-Checking eBook Content of Glencoe Chemistry Text Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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