- Assign oxidation numbers to all atoms in the equation.
- 2. Identify the atoms that are oxidized and the atoms that are reduced.
- Determine the change in oxidation number for the atoms that are oxidized and for the atoms that are reduced.
- Make the change in oxidation numbers equal in magnitude by adjusting coefficients in the equation.
- If necessary, use the conventional method to balance the remainder of the equation.

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Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill,2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change Chemistry John S. Phillips,Cheryl Wistrom,2000

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Electrons, Atoms, and Molecules in Inorganic Chemistry Joseph J. Stephanos, Anthony W. Addison, 2017-06-01 Electrons Atoms and Molecules in Inorganic Chemistry A Worked Examples Approach builds from fundamental units into molecules to provide the reader with a full understanding of inorganic chemistry concepts through worked examples and full color illustrations. The book uniquely discusses failures as well as research success stories. Worked problems include a variety of types of chemical and physical data illustrating the interdependence of issues. This text contains a bibliography providing access to important review articles and papers of relevance as well as summaries of leading articles and reviews at the end of each chapter so interested readers can readily consult the original literature Suitable as a professional reference for researchers in a variety of fields as well as course use and self study. The book offers valuable information to fill an important gap in the field Incorporates questions and answers to assist readers in understanding a variety of problem types Includes detailed explanations and developed practical approaches for solving real chemical problems Includes a range of example

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