

Name _____ Date _____ Class _____

CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	1. 4 molecules	2. O_2
1.00 mol	2.00 mol	2.00 mol	3. none	4. none
4.00 mol	6.00 mol	4.00 mol	5. NO	6. 2.00 mol O_2
6.00 mol	12.00 mol	6.00 mol	7. NO	8. 1.50 mol O_2
1.00 mol	6.00 mol	10.00 mol	9. O_2	10. 5.00 mol NO
6.000 mol	12.000 mol	16.000 mol	11. NO	12. 6.000 mol O_2
32.00 g	60.02 g	16.02 g	13. none	14. none
64.00 g	60.02 g	16.02 g	15. O_2	16. 30.12 g NO
64.00 g	120.04 g	32.04 g	17. O_2	18. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g $\text{O}_2 \times \frac{1 \text{ mol } \text{O}_2}{32.00 \text{ g } \text{O}_2} = 0.5125 \text{ mol } \text{O}_2$

20.00 g $\text{NO} \times \frac{1 \text{ mol } \text{NO}}{30.01 \text{ g } \text{NO}} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $\frac{0.6664 \text{ mol } \text{NO}}{0.5125 \text{ mol } \text{O}_2} = 1.302 \text{ mol } \text{NO}/1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of NO/O_2 is larger than the balanced equation mole ratio of NO/O_2 , there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times \frac{2 \text{ mol } \text{NO}}{1 \text{ mol } \text{O}_2} = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times \frac{30.01 \text{ g } \text{NO}}{1 \text{ mol } \text{NO}} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times \frac{46.01 \text{ g } \text{NO}_2}{1 \text{ mol } \text{NO}_2} = 19.56 \text{ g } \text{NO}_2$

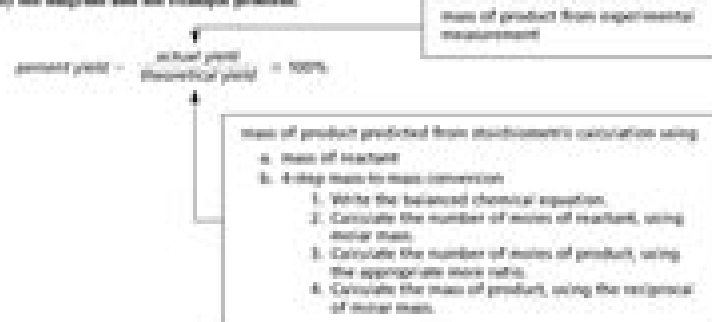
Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

Name _____ Date _____ Class _____

CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield unknown 7.42 g Ga_2O_3
2. mass of reactant: 7.00 g Ga correct
3. number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{1 \text{ mol } \text{Ga}}{69.72 \text{ g } \text{Ga}} = 0.100 \text{ mol } \text{Ga}$
4. number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}} = 0.200 \text{ mol } \text{Ga}_2\text{O}_3$
5. theoretical yield: $0.200 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{157.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 31.49 \text{ g } \text{Ga}_2\text{O}_3$ correct
6. percent yield: $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{31.49 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 23.6\%$

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Table of Contents Glencoe Chemistry Answer Key

1. Understanding the eBook Glencoe Chemistry Answer Key
 - The Rise of Digital Reading Glencoe Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glencoe Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Answer Key
 - Personalized Recommendations
 - Glencoe Chemistry Answer Key User Reviews and Ratings
 - Glencoe Chemistry Answer Key and Bestseller Lists
5. Accessing Glencoe Chemistry Answer Key Free and Paid eBooks
 - Glencoe Chemistry Answer Key Public Domain eBooks
 - Glencoe Chemistry Answer Key eBook Subscription Services
 - Glencoe Chemistry Answer Key Budget-Friendly Options

6. Navigating Glencoe Chemistry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Glencoe Chemistry Answer Key Compatibility with Devices
 - Glencoe Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glencoe Chemistry Answer Key
 - Highlighting and Note-Taking Glencoe Chemistry Answer Key
 - Interactive Elements Glencoe Chemistry Answer Key
8. Staying Engaged with Glencoe Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glencoe Chemistry Answer Key
9. Balancing eBooks and Physical Books Glencoe Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glencoe Chemistry Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Glencoe Chemistry Answer Key
 - Setting Reading Goals Glencoe Chemistry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Glencoe Chemistry Answer Key
 - Fact-Checking eBook Content of Glencoe Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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