

MENG 22100, Spring 2020

Homework Assignment 2

(Due 4/24/2020)

HW 2-1 What is activation energy? Describe two experimental methods that can be used to estimate it.

HW 2-2 Using expressions established by the collision theory, calculate the mean time between collisions experienced by an argon atom, diameter 0.29 nm, in (a) 200 kPa and (b) 200 Pa of argon gas at 298 K.

HW 2-3 Calculate the number of collisions per cubic centimeter per second between molecules A and B for a mixture containing 10 kPa of each at 300 K. The molecular diameters of A and B are 0.3 and 0.4 nm, and their atomic masses are 2 and 32 amu, respectively. Given that the experimental rate constant for the reaction is $1.07 \times 10^{-5} \text{ cm}^3 \text{ mol}^{-1} \text{ s}^{-1}$ at 300 K, and that the activation energy is 40 kJ mol^{-1} , (a) calculate the fraction of the collisions that result in a reaction and (b) determine the steric factor of the reaction. (Steric factor is defined as the ratio between the experimental frequency factor and that predicted by the collision theory.)

HW 2-4 For the reaction between atomic fluorine and molecular hydrogen at 300 K:



calculate the rate constant using collision theory and Polanyi equation. Additional useful information is provided as follows.

Reactants

Effective molecular diameter for H_2 : 2.5 Å

Effective atomic diameter for F: 1.7 Å

F: 19 amu

H_2 : 2 amu

Transition State for Reaction $\text{Br} + \text{H}_2 \rightarrow \text{HBr} + \text{H}$

$E_a = 552 \text{ kJ/mol}$

Linear configuration: $\text{Br} \cdots \text{H} \cdots \text{H}$

Distances: $\text{Br} \cdots \text{H}$ 1.62 Å; $\text{H} \cdots \text{H}$ 0.78 Å

Bond length for H_2 : 0.74 Å

Bond length for HBr: 1.41 Å

Heats of Formation

HF: -272.55 kJ/mol

HBr: -36.44 kJ/mol

H: 218 kJ/mol

F: 79.39 kJ/mol

Br: 111.9 kJ/mol

H_2 : 0 kJ/mol

Gizmo Geeks On Trial With Answers

Victor M. Corman



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Table of Contents Gizmo Geeks On Trial With Answers

1. Understanding the eBook Gizmo Geeks On Trial With Answers
 - The Rise of Digital Reading Gizmo Geeks On Trial With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmo Geeks On Trial With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmo Geeks On Trial With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmo Geeks On Trial With Answers
 - Personalized Recommendations
 - Gizmo Geeks On Trial With Answers User Reviews and Ratings
 - Gizmo Geeks On Trial With Answers and Bestseller Lists
5. Accessing Gizmo Geeks On Trial With Answers Free and Paid eBooks
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 - Gizmo Geeks On Trial With Answers eBook Subscription Services
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 - ePub, PDF, MOBI, and More
 - Gizmo Geeks On Trial With Answers Compatibility with Devices
 - Gizmo Geeks On Trial With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmo Geeks On Trial With Answers
 - Highlighting and Note-Taking Gizmo Geeks On Trial With Answers
 - Interactive Elements Gizmo Geeks On Trial With Answers
8. Staying Engaged with Gizmo Geeks On Trial With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmo Geeks On Trial With Answers
9. Balancing eBooks and Physical Books Gizmo Geeks On Trial With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmo Geeks On Trial With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gizmo Geeks On Trial With Answers
 - Setting Reading Goals Gizmo Geeks On Trial With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gizmo Geeks On Trial With Answers
 - Fact-Checking eBook Content of Gizmo Geeks On Trial With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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