



## Student Exploration: Limiting Reactants

**Learning Objectives:** chemical equation, chemical formula, chemical reactions, coefficient, limiting reactant, molecules, product, substance, substance

**Prior Knowledge:** Speedometer (distance and time) measuring the distance of a runner's race.

Imagine you are a runner who has just finished a 100-meter race. You are standing at the finish line and you are looking at the clock. You are looking at the clock and you are looking at the clock. You are looking at the clock and you are looking at the clock.

1. How many molecules of hydrogen gas ( $H_2$ ) are there?

2. How many molecules of oxygen gas ( $O_2$ ) are there?

3. How many molecules of water ( $H_2O$ ) are there?

**Answer:** 100 molecules

Now we are going to look at the reaction between hydrogen and oxygen. The reaction is shown in the diagram below. The reaction is shown in the diagram below. The reaction is shown in the diagram below.

Hydrogen is shown as  $H_2$  molecules, and oxygen is shown as  $O_2$  molecules. The reaction is shown in the diagram below. The reaction is shown in the diagram below. The reaction is shown in the diagram below.

Hydrogen, which is  $H_2$ , is shown as  $H_2$  molecules. The reaction is shown in the diagram below. The reaction is shown in the diagram below. The reaction is shown in the diagram below.

4. How many molecules of hydrogen gas ( $H_2$ ) are there?

5. How many molecules of oxygen gas ( $O_2$ ) are there?

6. How many molecules of water ( $H_2O$ ) are there?

7. How many molecules of hydrogen gas ( $H_2$ ) are there?

8. How many molecules of oxygen gas ( $O_2$ ) are there?



# Gizmo Explore Learning Student Exploration Limiting Reactants Answers

**S Marginson**



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