



Roll No.

Answer Sheet No.

23

Sig. of Candidate.

Sig. of Invigilator.

CHEMISTRY SSC-I
SECTION – A (Marks 12)

Time allowed: 20 Minutes

NOTE: Section-A is compulsory. All parts of this section are to be answered on the question paper itself. It should be completed in the first 20 minutes and handed over to the Centre Superintendent. Deleting/overwriting is not allowed. Do not use lead pencil.

Q. 1 Circle the correct option i.e. A / B / C / D. Each part carries one mark.

- (i) If the molar mass of a compound is 90 g/mole, the number of its moles containing 3.011×10^{23} molecules will be:
A. 1.5 B. 3.0 C. 0.5 D. 4.5
- (ii) The correct electronic configuration of $^{24}_{12}\text{Mg}$ is:
A. $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^4$ B. $1s^2, 2s^2, 2p^6, 3s^2$
C. $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^1, 3d^3$ D. $1s^2, 2s^2, 2p^6, 3s^4$
- (iii) If you move from left to right in a period of periodic table, which of the following will **NOT** change?
A. Atomic radius B. Atomic mass C. Shielding effect D. Nuclear charge
- (iv) The bond which is formed by complete transfer of electrons from one atom to another is called:
A. Covalent bond B. Hydrogen bond
C. Co-ordinate bond D. Ionic bond
- (v) The melting point and boiling point of water are higher due to:
A. Its covalent nature B. Hydrogen bonding
C. Its higher molecular mass D. Its higher density
- (vi) The equation $\frac{V}{T} = \text{constant}$ is the mathematical expression of:
A. Charles's law B. Avogadro's law
C. Boyle's law D. Law of diffusion
- (vii) The direct conversion of a solid into vapours is called:
A. Evaporation B. Melting C. Boiling D. Sublimation
- (viii) An aqueous solution can dissolve 25 g of a solid at 25°C . If someone dissolves 20 g of this solid in water at 25°C then this solution will be called:
A. Supersaturated solution B. Unsaturated solution
C. Saturated solution D. Solution of liquid in liquid
- (ix) A chemist dissolves 15 g of NaCl in 85g of H_2O . This solution is:
A. 85% m/m NaCl solution B. 15% m/v NaCl solution
C. 15% v/v NaCl solution D. 15% m/m NaCl solution
- (x) Gain of electrons and loss of oxygen is called:
A. Oxidation B. Reduction C. Redox reaction D. Neutralization reaction
- (xi) Which of the following is the reaction taking place at anode?
A. $2\text{H}_{2(\text{g})} + \text{O}_{2(\text{g})} \rightarrow 2\text{H}_2\text{O}_{(\text{l})}$ B. $\text{Zn}^{+2}_{(\text{aq})} + 2\text{e}^- \rightarrow \text{Zn}_{(\text{s})}$
C. $\text{Cr}_{(\text{s})} \rightarrow \text{Cr}^{+3}_{(\text{aq})} + 3\text{e}^-$ D. $\text{Pb}_{(\text{s})} + \text{Cl}_{2(\text{g})} \rightarrow \text{PbCl}_{2(\text{s})}$
- (xii) Which of the following oxides is the **Most** basic oxide?
A. SrO B. BaO C. CaO D. MgO

For Examiner's use only:

Total Marks:

12

Marks Obtained:

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E Durkheim



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