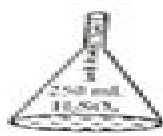


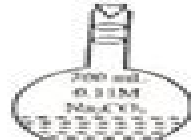
অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পরীক্ষার কপূর, এখানে কোর পরীক্ষা, আরও কলেজের শিক্ষার্থীদের কলেজসমূহের পরীক্ষা এবং মতের টেমের পরীক্ষা কপূর উত্তর অধ্যায়ভিত্তিক সেবা করা হয়েছে। এগুলোর অনুশীলন করতে কৃপা এ অধ্যায়টি থেকে লেখকের কৃপার উপস্থাপন করা হয়েছে।



পাত্র-A



পাত্র-B

(সং. ক্র. ১১-১ এর পৃষ্ঠা ১)

- ক. মোলার প্রবণ কী? ১
 - খ. "কোনো পানীয় লবণের প্রাকৃতিক উপস্থিতির সাথে সাথে এক পর্যায়ে কৃষ্ণ না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো। ২
 - গ. A-পাত্রে বিদ্যমান এলিভের ভর কত নির্ণয় করো। ৩
 - ঘ. দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অম্লীয় বা ক্ষারীয় হবে— বিশ্লেষণ করো। ৪
- ১ নং প্রশ্নের উত্তর

1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কোনো পানীয় লবণের প্রাকৃতিক উপস্থিতির সাথে সাথে এক পর্যায়ে কৃষ্ণ না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা কৃষ্ণের সাথে সাথে প্রাকৃতিক কৃষ্ণ পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কোনো পানীয় লবণ নিষ্ক্রিয় হয়। তখন প্রাকৃতিক হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরূপ বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা কৃষ্ণজনিত কারণে হ্রাসের হ্রাস ঘটে।

আমরা জানি,
 $w = MV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
 অতএব, দ্রবীভূত H_2SO_4 এর ভর = 12.25 g

উদাহরণের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
 দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

পাত্র-A: H_2SO_4 এর মোল সংখ্যা = $\left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা = $\left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.0275 \text{ mol}$

সমস্ত বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 সন্ধিকরণ ঘটা পাই, 1 mol H_2SO_4 = 1 mol Na_2CO_3
 1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.0275 mol " " " " " " " " " "
 $(0.0275) \text{ mol}$ " " " " " " " " " "
 $= 0.0275 \text{ mol}$ " " " " " " " " " "

অবশিষ্ট H_2SO_4 এর পরিমাণ = $(0.125 - 0.0275) \text{ mol}$
 $= 0.0975 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.0275 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

একজন ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর তরুণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রবণ তৈরি করে।

- ক. লোভ কাকে বলে? ১
- খ. Cl_2 অপোলার— ব্যাখ্যা করো। ২
- গ. উদাহরণের এলিভটির মোলার ঘনমাত্রা নির্ণয় করো। ৩
- ঘ. উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো। ৪

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে লোভ বলে।

সমযোজী বন্ধনে আকর্ষণ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা কম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন ক্রিয়াকারী ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl_2 এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিপরীত সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এরূপ থাকে।

উদাহরণে বর্ণিত এলিভটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL দ্রবণে HCl এর ভর 2g

∴ 1000 mL " " " " " " " " " "
 $= \left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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