

**Answer: Soluble Precipitates**

**Answer:** In this activity we will address the question: **How do I construct formulas of ionic compounds that precipitate?**

**Answer:**

1.  $Ca^{2+}$  (aq) +  $CO_3^{2-}$  (aq) →  $CaCO_3$  (s)

2.  $Ca^{2+}$  (aq) +  $SO_4^{2-}$  (aq) →  $CaSO_4$  (s)

3.  $Ca^{2+}$  (aq) +  $PO_4^{3-}$  (aq) →  $Ca_3(PO_4)_2$  (s)

4.  $Ca^{2+}$  (aq) +  $S^{2-}$  (aq) →  $CaS$  (s)

5.  $Ca^{2+}$  (aq) +  $OH^{-}$  (aq) →  $Ca(OH)_2$  (s)

6.  $Ca^{2+}$  (aq) +  $C_2O_4^{2-}$  (aq) →  $CaC_2O_4$  (s)

7.  $Ca^{2+}$  (aq) +  $F^{-}$  (aq) →  $CaF_2$  (s)

8.  $Ca^{2+}$  (aq) +  $SiO_3^{2-}$  (aq) →  $CaSiO_3$  (s)

9.  $Ca^{2+}$  (aq) +  $CO_3^{2-}$  (aq) →  $CaCO_3$  (s)

10.  $Ca^{2+}$  (aq) +  $SO_4^{2-}$  (aq) →  $CaSO_4$  (s)

11.  $Ca^{2+}$  (aq) +  $PO_4^{3-}$  (aq) →  $Ca_3(PO_4)_2$  (s)

12.  $Ca^{2+}$  (aq) +  $S^{2-}$  (aq) →  $CaS$  (s)

13.  $Ca^{2+}$  (aq) +  $OH^{-}$  (aq) →  $Ca(OH)_2$  (s)

14.  $Ca^{2+}$  (aq) +  $C_2O_4^{2-}$  (aq) →  $CaC_2O_4$  (s)

15.  $Ca^{2+}$  (aq) +  $F^{-}$  (aq) →  $CaF_2$  (s)

16.  $Ca^{2+}$  (aq) +  $SiO_3^{2-}$  (aq) →  $CaSiO_3$  (s)

17. **What is the coefficient of a balanced equation?**

18. **What is the coefficient of a balanced equation?** **What is the coefficient of a balanced equation?**

19. **What is the coefficient of a balanced equation?** **What is the coefficient of a balanced equation?**

20. **What is the coefficient of a balanced equation?** **What is the coefficient of a balanced equation?**

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**Y Pai**





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