



# Hampton Bay Transitional Lighting Manual

**RD Boyd**



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subscript a for acid

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web write down the equilibrium equation you are investigating using the information provided by the introduction within the virtual lab activity equilibrium reaction  $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{HCO}_3^-$  select each control test tube to compare the concentrations of

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web equilibrium moles  $\text{HSCN}$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{HSCN}$  knowing the number of moles of  $\text{Fe}^{3+}$  and  $\text{HSCN}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $\text{mol Fe}^{3+}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{Fe}^{3+}$   $8.50 \times 10^{-4} \text{ M}$

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equivalent to the ratio of the  $k_{eq}$  value of a covalent bond to that of a hydrogen bond

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