

Answer Key

60 + 4 ext

Acids and Bases

How do acids and bases behave in water?

Why?

Acids and bases play an important role in our lives. Numerous biological processes, industrial applications, and even environmental problems are a function of the acidity or basicity (alkalinity) of aqueous solutions. It is therefore important to understand what makes a substance behave as an acid or a base when dissolved in water. In this activity, we will explore the physical and chemical properties of acids and bases.

Model 1 – Arrhenius Acids and Bases

Common Name for Aqueous Solution	Chemical Formula	Found in...	Tastes...	Turns Litmus Paper...	Conducts Electricity?	Acid or Base
Acetic acid	$\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$	Vinegar	Sour	Red	Yes	Acid
Benzoic acid	$\text{HC}_7\text{H}_5\text{O}_2(\text{aq})$	Food preservative	Sour	Red	Yes	Acid
Phosphoric acid	$\text{H}_3\text{PO}_4(\text{aq})$	Soda pop	Sour	Red	Yes	Acid
Hydrochloric acid	$\text{HCl}(\text{aq})$	Stomach acid	Sour	Red	Yes	Acid
Citric acid	$\text{H}_3\text{C}_6\text{H}_7\text{O}_7(\text{aq})$	Citrus fruits	Sour	Red	Yes	Acid
Ascorbic acid	$\text{H}_2\text{C}_6\text{H}_8\text{O}_6(\text{aq})$	Vitamin C	Sour	Red	Yes	Acid
Magnesium hydroxide	$\text{Mg}(\text{OH})_2(\text{aq})$	Milk of magnesia	Bitter	Blue	Yes	Base
Aluminum hydroxide	$\text{Al}(\text{OH})_3(\text{aq})$	Antacids	Bitter	Blue	Yes	Base
Barium hydroxide	$\text{Ba}(\text{OH})_2(\text{aq})$	Lubricants	POISON	Blue	Yes	Base
Sodium hydroxide	$\text{NaOH}(\text{aq})$	Drain cleaner	POISON	Blue	Yes	Base

1. Refer to Model 1.

1 a. What is the common chemical name for Vitamin C?

Ascorbic acid

1 b. Is Vitamin C classified as an acid or a base?

Acid

3 2. Examine the properties of the Arrhenius acids in Model 1. List three properties that all Arrhenius acids have in common.

sour, turn litmus paper red, conduct electricity

1 3. Examine the chemical formulas for the Arrhenius acids in Model 1. What feature do all the Arrhenius acid chemical formulas have in common?

H at the beginning

2 4. Examine the properties of the Arrhenius bases in Model 1. List two properties that all Arrhenius bases have in common.

bitter/poison, turn litmus paper blue, conduct electricity

Gpb Acids Bases Answers

M Tight



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competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it
will assess your knowledge and prepare you for competitive exams quizzes trivia and more *Acids and Bases* Brian G.
Cox,2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their
behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and
equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of
pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of
substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in
non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between
solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid
base equilibria including definitions and relationships between solution pH and species distribution the influence of
molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the
magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in
hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining
dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in
three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic
polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and
tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents

and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described Proton Chemistry Robert A. Richardson, 1976 Alberta Authorized Resource for grade 12 ca 1980 1994

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Table of Contents Gpb Acids Bases Answers

1. Understanding the eBook Gpb Acids Bases Answers
 - The Rise of Digital Reading Gpb Acids Bases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gpb Acids Bases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpb Acids Bases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gpb Acids Bases Answers
 - Personalized Recommendations

- Gpb Acids Bases Answers User Reviews and Ratings
- Gpb Acids Bases Answers and Bestseller Lists
- 5. Accessing Gpb Acids Bases Answers Free and Paid eBooks
 - Gpb Acids Bases Answers Public Domain eBooks
 - Gpb Acids Bases Answers eBook Subscription Services
 - Gpb Acids Bases Answers Budget-Friendly Options
- 6. Navigating Gpb Acids Bases Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gpb Acids Bases Answers Compatibility with Devices
 - Gpb Acids Bases Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Acids Bases Answers
 - Highlighting and Note-Taking Gpb Acids Bases Answers
 - Interactive Elements Gpb Acids Bases Answers
- 8. Staying Engaged with Gpb Acids Bases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Acids Bases Answers
- 9. Balancing eBooks and Physical Books Gpb Acids Bases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Acids Bases Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gpb Acids Bases Answers
 - Setting Reading Goals Gpb Acids Bases Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gpb Acids Bases Answers
 - Fact-Checking eBook Content of Gpb Acids Bases Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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