

HL Paper 2

Chloroethene, $\text{C}_2\text{H}_3\text{Cl}$, is an important organic compound used to manufacture the polymer poly(chloroethene).

- d.i. State an equation for the reaction of ethanoic acid with water. [1]
- d.ii. Calculate the pH of $0.200 \text{ mol dm}^{-3}$ ethanoic acid ($\text{p}K_{\text{a}} = 4.76$). [3]
- e. Determine the pH of a solution formed from adding 50.0 cm^3 of 1.00 mol dm^{-3} ethanoic acid, $\text{CH}_3\text{COOH}(\text{aq})$, to 50.0 cm^3 of $0.500 \text{ mol dm}^{-3}$ sodium hydroxide, $\text{NaOH}(\text{aq})$. [4]
- f. (If acid added) $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$. [2]
(If alkali added) $\text{CH}_3\text{COOH} + \text{OH}^- \rightarrow \text{CH}_3\text{COO}^- + \text{H}_2\text{O}$.
Explanation marks cannot be awarded without equations.

Water is an important substance that is abundant on the Earth's surface.

Buffer solutions resist small changes in pH. A phosphate buffer can be made by dissolving NaH_2PO_4 and Na_2HPO_4 in water, in which NaH_2PO_4 produces the acidic ion and Na_2HPO_4 produces the conjugate base ion.

A 0.10 mol dm^{-3} ammonia solution is placed in a flask and titrated with a 0.10 mol dm^{-3} hydrochloric acid solution.

- a. (i) State the expression for the ionic product constant of water, K_{w} . [1]
(ii) Explain why even a very acidic aqueous solution still has some OH^- ions present in it.
(iii) State and explain the effect of increasing temperature on the value of K_{w} given that the ionization of water is an endothermic process.
(iv) State and explain the effect of increasing temperature on the pH of water.
- b. (i) Deduce the acid and conjugate base ions that make up the phosphate buffer and state the ionic equation that represents the phosphate buffer. [7]
(ii) Describe how the phosphate buffer minimizes the effect of the addition of a strong base, $\text{OH}^-(\text{aq})$, to the buffer. Illustrate your answer with an ionic equation.
(iii) Describe how the phosphate buffer minimizes the effect of the addition of a strong acid, $\text{H}^+(\text{aq})$, to the buffer. Illustrate your answer with an ionic equation.
- c. (i) Explain why the pH of the ammonia solution is less than 12. [11]
(ii) Estimate the pH at the equivalence point for the titration of hydrochloric acid with ammonia and explain your reasoning.
(iii) State the equation for the reaction of ammonia with water and write the K_{b} expression for $\text{NH}_3(\text{aq})$.

M13 Chemistry H1 Paper 2

Judd E. Hollander



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