

## Quantitative analysis

Read through the whole method before starting any practical work. Where appropriate, prepare a table for your results in the space provided.

Show your working and appropriate significant figures in the final answer to **each** step of your calculations.

- 1 In this experiment you will carry out a titration to determine the relative formula mass,  $M_r$ , of a monoprotic acid, HX.

**FA 1** is HX, the monoprotic acid.

**FA 2** is  $0.0450 \text{ mol dm}^{-3}$  sodium carbonate,  $\text{Na}_2\text{CO}_3$ .  
methyl orange indicator

### (a) Method

#### Preparing a solution of FA 1

- Weigh the empty  $250 \text{ cm}^3$  beaker. Record the mass.
- Transfer all the **FA 1** into the beaker.
- Weigh the beaker and **FA 1**. Record the mass.
- Calculate and record the mass of **FA 1** used.
- Add approximately  $100 \text{ cm}^3$  of distilled water to the **FA 1** in the beaker.
- Stir the mixture with a glass rod until all the **FA 1** has dissolved.
- Transfer this solution into the  $250 \text{ cm}^3$  volumetric flask.
- Wash the beaker with distilled water and transfer the washings to the volumetric flask.
- Rinse the glass rod with distilled water and transfer the washings to the volumetric flask.
- Make up the solution in the volumetric flask to the mark using distilled water.
- Shake the flask thoroughly.
- This solution of HX is **FA 3**. Label the flask **FA 3**.

### Results

#### Titration

- Fill the burette with **FA 3**.
- Pipette  $25.0 \text{ cm}^3$  of **FA 2** into a conical flask.
- Add several drops of methyl orange indicator.
- Perform a rough titration and record your burette readings in the space below.

The rough titre is .....  $\text{cm}^3$ .

# June 2013 Chemistry 9701 Paper 35 Question

**DJ Losen**



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