

সেট : ০২

পদার্থবিজ্ঞান (সৃজনশীল)
[২০২১ সালের সিলেবাস অনুযায়ী]
প্রথম পত্র

বিষয় কোড :

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সময়—২ ঘণ্টা ৩৫ মিনিট

পূর্ণমান—৫০

[দ্রষ্টব্য : ডান পাশের সংখ্যা প্রশ্নের পূর্ণমান জ্ঞাপক। প্রদত্ত উদ্দীপকগুলো মনোযোগসহকারে পড় এবং সংশ্লিষ্ট প্রশ্নগুলোর যথাযথ উত্তর দাও। যে কোনো পাঁচটি প্রশ্নের উত্তর দিতে হবে।]

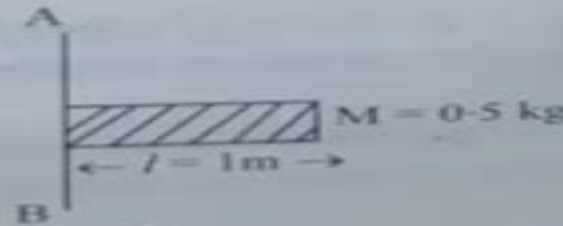
নম্বর



10km প্রস্থবিশিষ্ট একটি নদীতে স্রোতের বেগ 5kmh^{-1} । প্রথম মাঝি 10kmh^{-1} বেগে স্রোতের সাথে α কোণে এবং ২য় মাঝি 10kmh^{-1} বেগে স্রোতের সাথে লম্বভাবে নদী পার হতে যাত্রা করল।

- বলের ঘাত কী?
- ঘর্ষণ বল কি অসংরক্ষণশীল বল? ব্যাখ্যা কর।
- α কোণের মান কত হলে ১ম মাঝি সোজাসুজি নদীর অপর পাড়ে পৌঁছাবে?
- ২য় কোণে মাঝি নদীর অপর পাড়ে আগে পৌঁছাতে পারবে? গাণিতিক বিশ্লেষণসহ মতামত দাও।

২। একটি আয়তাকার দণ্ড যার দৈর্ঘ্য $l = 1\text{m}$ ও ভর $M = 0.5\text{ kg}$.



চিত্র

- গ্র্যাভিয়েন্ট কাকে বলে?
- ভেক্টরের ডাইরেকশনের বৈশিষ্ট্য কী কী?
- দণ্ডটির জড়তার আয়তক নির্ণয় কর।
- যদি দণ্ডটিকে 1m ব্যাসের পাকলা চাকতিতে পরিণত করা হয় তবে জড়তার আয়তকের কোনো পরিবর্তন হবে কি? গাণিতিক বিশ্লেষণ কর।

[পর পৃষ্ঠা দ্রষ্টব্য]

Hsc Physics Question Paper 22

RJ Alexander

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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web 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ 80 0 grams of iodine v oxide I_2O_5 reacts with 28 0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0 160 moles of iodine I_2 was produced i what mass of iodine was produced

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ c determine the percent yield yield actual $\times 100$ $260.0 \text{ g} \times 100 / 86.01$ theoretical 302.3 g

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web 7 Na $2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H 2 are left over 13

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