

ক) ভারী হ্যালোজেনের সাথে বিজোড় মতখ্যক হালকা হ্যালোজেন যুক্ত হয়ে আন্তঃ হ্যালোজেন যৌগ গঠন করে।

ক) ~~হালকা~~ হ্যালোজেন: হ্যালোজেন নয় কিন্তু হ্যালোজেনের মতোই রাসায়নিক বিক্রিয়া করে। যেমন: CN^- , CNS^- ,

ক) ট্রান্স-ইউরেনিয়াম মৌল: (৯৩-১০৩) পারমানবিক মতখ্যাবিশিষ্ট মৌলগুলোকে ট্রান্স-ইউরেনিয়াম মৌল বলে।

ক) কুপের মৌল: (৯৩-১১৪) পারমানবিক মতখ্যাবিশিষ্ট মৌলগুলোকে কুপের মৌল বলে।

ক) প্রকৃতিতে পাওয়া মৌল ৯০ টি।

ক) (১-৯২) পারমানবিক মতখ্যাবিশিষ্ট মৌলগুলোর মধ্যে দুটি মৌল প্রকৃতিতে পাওয়া যায় না। Te (৫৩) একমাত্র Po (৮৪)।

ক) ১ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^{1-2}$

০ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2$

১ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^1$

২ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^2$

৩ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^3$

৪ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^4$

৫ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^5$

৬ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^6$

৭ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^7$

৮ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^8$

৯ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^9$

১০ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}$

১১ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^1$

১২ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^2$

১৩ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^3$

১৪ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^4$

১৫ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^5$

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২১ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^{10}(n-1)g^1$

২২ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^{10}(n-1)g^2$

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২৯ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^{10}(n-1)g^9$

৩০ ব্রক মৌলের বহিঃস্থ ইলেকট্রন কাঠামো $\rightarrow ns^2(n-1)d^{10}(n-1)f^{10}(n-1)g^{10}$

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