

SECTION - I

Q. 1. Select and write the most appropriate answer from the given alternatives for each sub-question : [7]

- (i) The hybridisation of phosphorus in phosphorus pentachloride is-
 (a) dsp^3 (b) sp^3d (c) d^2sp^3 (d) sp^3d^2
- (ii) The rate constant for a first order reaction is 100 S^{-1} . The time required for completion of 50% of reaction is-
 (a) 0.0693 milliseconds (b) 0.693 milliseconds
 (c) 6.93 milliseconds (d) 69.3 milliseconds
- (iii) Silica is added to roasted copper ore during smelting process to remove -
 (a) ferrous sulphide (b) ferrous oxide (c) cuprous sulphide (d) cuprous oxide
- (iv) 96500 coulombs correspond to the charge on how many electrons ?
 (a) 1.6×10^{19} (b) 6.022×10^{20} (c) 6.022×10^{23} (d) 6.022×10^{24}
- (v) For the reaction : $\text{Cl}_{2(g)} \longrightarrow 2\text{Cl}_{(g)}$ -
 (a) ΔH is positive, ΔS is positive (b) ΔH is positive, ΔS is negative
 (c) ΔH is negative, ΔS is negative (d) ΔH is negative, ΔS is positive
- (vi) The substance 'X', when dissolved in solvent water gave molar mass corresponding to the molecular formula ' X_3 '. The van't Hoff factor (i) is -
 (a) 3 (b) 0.33 (c) 1.3 (d) 1
- (vii) The relation $a \neq b \neq c$ and $\alpha \neq \beta \neq \gamma \neq 90^\circ$ represents which crystal system ?
 (a) Orthorhombic (b) Tetragonal (c) Triclinic (d) Monoclinic

Q. 2. Answer any SIX of the following :

[12]

- (i) Define ebullioscopic constant. Write its unit.
- (ii) State : (a) Second law of thermodynamics in terms of entropy.
 (b) Third law of thermodynamics.
- (iii) What is the hybridisation of carbon atom in diamond and graphite ?
 0.1 mole of Buckminster fullerene contains how many Kg of carbon ?
 [Atomic mass of carbon = 12]
- (iv) Draw structures of (a) Chlorine trifluoride (b) Chlorine pentafluoride.

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Roman Wölfel

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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