

Chemistry 12
Worksheet 2-2
LeChatelier's Principle Name _____

1. In order to decide what effect a **change in total pressure** will have on an equilibrium system with gases, what is the first thing you should do when given the balanced equation?

2. Predict which way the following equilibrium systems will shift when the **total pressure** is **increased**. (NOTE: Some may have no shift)



3. Which way will the following equilibrium shift if the **total pressure** on the system is **decreased**?



4. Explain why a flask filled with $NO_{2(g)}$ and $N_2O_{4(g)}$ will get **darker** when heated. Use the equation: $N_2O_{4(g)} + \text{heat} \rightleftharpoons 2NO_{2(g)}$

5. State **Le Chatelier's Principle**.

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