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# Hamilton International H74592433 Watches Owners Manual

**David Baud**



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid  $\text{H}_2\text{CO}_3$  and its conjugate base the bicarbonate ion  $\text{HCO}_3^-$

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arrhenius's definition of acids and bases the earliest definition of acids and bases is arrhenius's definition which states that an acid is a substance that forms hydrogen ions  $H^+$  when dissolved in water and a base is a substance that forms hydroxide ions  $OH^-$  when dissolved in water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is

$$\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$$

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