

1. How many moles of **atoms** are there in:

- (a) 64 g of copper
- (b) 28 g of iron

2. How many moles of **molecules** are there in:

- (a) 4.5 g of H_2O
- (b) 1.9 g of fluorine gas.

3. What is the mass of:

- (a) 3 moles of silver
- (b) 0.01 mole of V_2O_5

4. How many moles of **ions** are present in:

- (a) 3 mole of KBr
- (b) 0.02 moles of MgBr_2

5. How many moles of **atoms** are present in:

- (a) 0.4 moles of chlorine gas
- (b) 1 mole of carbon dioxide

6. What is the empirical formula for a compound with molecular formula $\text{C}_6\text{H}_6\text{N}_2\text{O}_4$?

7. Copy and balance the following equation:



8. A hydrocarbon has the empirical formula CH_2 and a molar mass of 70 g/mol. What is its molecular formula?

9. A compound has the following percentage composition:

K: 38.6 % N: 13.9 % O: 47.5 %

Work out the empirical formula of the compound.

10. Calculate the percentage by mass of each element in NH_4NO_3 .

11. Below are the results from an experiment carried out to find the formula of an oxide of copper. In this experiment some copper metal was weighed before and after heating to constant mass in a crucible.

Mass of crucible	= 15.6 g
Mass of crucible + copper	= 18.8 g
Mass of crucible + copper oxide	= 19.2 g

- (a) What is the mass of copper used?
- (b) What is the mass of copper oxide produced?
- (c) What is the mass of oxygen that combined with copper?
- (d) Using your answers to parts (a) and (c), calculate the empirical formula of the copper oxide (a table will be useful).

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Table of Contents Grade 10 Chemistry Practice Test With Memo

1. Understanding the eBook Grade 10 Chemistry Practice Test With Memo
 - The Rise of Digital Reading Grade 10 Chemistry Practice Test With Memo
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 10 Chemistry Practice Test With Memo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 10 Chemistry Practice Test With Memo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 10 Chemistry Practice Test With Memo
 - Personalized Recommendations
 - Grade 10 Chemistry Practice Test With Memo User Reviews and Ratings
 - Grade 10 Chemistry Practice Test With Memo and Bestseller Lists
5. Accessing Grade 10 Chemistry Practice Test With Memo Free and Paid eBooks
 - Grade 10 Chemistry Practice Test With Memo Public Domain eBooks
 - Grade 10 Chemistry Practice Test With Memo eBook Subscription Services
 - Grade 10 Chemistry Practice Test With Memo Budget-Friendly Options

6. Navigating Grade 10 Chemistry Practice Test With Memo eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 10 Chemistry Practice Test With Memo Compatibility with Devices
 - Grade 10 Chemistry Practice Test With Memo Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 10 Chemistry Practice Test With Memo
 - Highlighting and Note-Taking Grade 10 Chemistry Practice Test With Memo
 - Interactive Elements Grade 10 Chemistry Practice Test With Memo
8. Staying Engaged with Grade 10 Chemistry Practice Test With Memo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 10 Chemistry Practice Test With Memo
9. Balancing eBooks and Physical Books Grade 10 Chemistry Practice Test With Memo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 10 Chemistry Practice Test With Memo
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 10 Chemistry Practice Test With Memo
 - Setting Reading Goals Grade 10 Chemistry Practice Test With Memo
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 10 Chemistry Practice Test With Memo
 - Fact-Checking eBook Content of Grade 10 Chemistry Practice Test With Memo
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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