

ex: The pH of blood is 7.41. Find the pOH,  $[H^+]$ , and  $[OH^-]$ .

$$pH + pOH = 14.00$$

$$7.41 + pOH = 14.00$$

$$pOH = 6.59$$



$$pOH = -\log [OH^-]$$

$$-pOH = \log [OH^-]$$

$$10^{-pOH} = 10^{\log [OH^-]}$$

$$10^{-pOH} = [OH^-]$$

$$pH = -\log [H^+]$$

$$-pH = \log [H^+]$$

$$10^{-pH} = 10^{\log [H^+]}$$

$$10^{-pH} = [H^+]$$

"10 to the log of something equals that something."

$$10^{-7.41} = [H^+] = 3.9 \times 10^{-8} M$$

→ punch in to calculator to get ↗

$$10^{-6.59} = [OH^-] = 2.6 \times 10^{-7} M$$

### \* Calculating the pH of Strong Acid Solutions.

↳  $HCl$ ,  $HNO_3$ ,  $H_2SO_4$ ,  $HClO_4$

- it is essential to know the major components (species) in the solution, and to know the minor species that can be ignored.

ex:  $[HNO_3]$  → a very strong acid, and strong acids dissociate completely in solution.

↳ so, a 1.0M  $HNO_3$  solution contains no  $HNO_3$ !! Instead it contains  $H^+$ ,  $NO_3^-$ , and  $H_2O$ .

ex: Calculate the pH of a 0.10M  $HCl$  solution.



so 0.10M  $H^+$

↳  $H^+$ ,  $Cl^-$ ,  $H_2O$  (no " $HCl$ ")

↳ its " $H^+$ " contribution is negligible.  
( $1.0 \times 10^{-7} M H^+$ )

$$pH = -\log [H^+] = -\log (0.10) \Rightarrow pH = 1.00$$

# Gpb Chemistry Answers Solutions Key

**D Kirk**



## **Gpb Chemistry Answers Solutions Key:**

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