

UNIT REVIEW - BONDING

Intramolecular Forces

1. Define electronegativity

- the tendency of atoms to attract e^-

2. Electronegativity increases as you go across the periodic table, and decreases as you go down a group.

3. Explain why atoms with smaller atomic radii have stronger electronegativities

- smaller atoms means e^- 's are closer to the nucleus, which has the protons
- if you decrease the distance between protons + e^- 's, you increase attractive forces.

4. Define intramolecular force

- forces that hold atoms together: ionic, covalent, polar covalent bond

5. Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	Ionic
- e^- are shared equally - between 2 non-metals - difference in electronegativity is ≤ 0.2 - atoms are neutral	- e^- are shared unequally - between 2 non-metals - difference in electronegativity is $0.3 - 1.6$ - atoms are partially charged	- e^- are transferred - between metal + non metal - difference in electronegativity is ≥ 1.7 - atoms have a charge

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal sharing of electrons between 2 atoms

An ionic charge results from a transfer of electrons from 1 atom to another

Flocabulary Unit 10 Changes Answer Key

Arturo Cuomo



Flocabulary Unit 10 Changes Answer Key:

Flocabulary Unit 10 Changes Answer Key Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Flocabulary Unit 10 Changes Answer Key**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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