



# Stoichiometry and Solubility

# **Mole Ratios and Chemical Formulas**

#### Introduction

Disable replacement reactions are generally unwadered to be inteverable. The formation of an involuble precipitate provides a driving force that makes the reaction proceed in one direction only. The purpose of this demonstration is to find the optimizers made ratio for the formation of a precipitate in a double replacement reaction and to use this information to precipitate the chemical formals of the precipitate.

# Concepts

- Sanichlammerry

- Made ratio

. Disable replacement reaction

Sofial sility rather

#### Materials

Copper(III chloride, CarCl<sub>2</sub>, 0.05 M, 250 ml., InsofIII) nitrate, Fe(NO<sub>2</sub>)<sub>2</sub>, 0.1 M, 250 ml., Stalium hydrostie, NaOH, 0.1 M, 250 ml., Stalium phosphate, NaJPO<sub>2</sub>, 0.05 M, 250 ml., Gradusted cylinders, 50-ml., 2 Gradusted cylinders, 100-ml., 7 Stirring train, large, 2

# Safety Precautions

Copper(IT) chloride, iron(ITI) intrate, exclusive hydroxide and exclusive phosphate educations are skin and eye irritants and are shightly trate by ingretion. Arosid contact with eyes and skin. West chemical ephasis gaggles, chemical-resistant gloves and a chemical-resistant appear. Please arounds correct a surround adject information.

#### Procedure

#### Part A. Reaction of Iron(III) Nitrate with Sodium Hydroxide

- Label seven 100-mL graduated cylinders 1–7.
- Using a clean, 50-ml, graduated cylinder, add the appropriate volume of montHH: nitrate solution to each graduated cylinder as shown in Table 1.
- Use a second 50-miligradu and cylinder to add the appropriate volume of sodium hydroxide solution to each graduated cylinder, as shown in Table 1.

#### Timbile 1.

Lanc L.									
Cylinder	1	- 2	3.	3.3	300	160	7-		
Fc(NO <sub>2</sub> ), 0.1 M, mL	100	1.5	2x0	30	40	45	50		
NaOII, 0.1 M, naL	50	4.5	-10	30	20	1.5	140		
Fe' :OH Mole Racio	1:5	1:3:	1.2	1:1	2:1	361	5:1		

- 4. The a large stirring real to thermughly note the reactants. Otherwe the signs of chemical reaction in each cylinder (Mining the yel-line-rating) adultion of inordIII) retrate with the colorless exclusion hydroxice solution gives a materialized precipitate and a pule yellow supernature.)
- Let the reaction ministers six antistaring for at least 10 minutes to allow the precipitates to settle. During this time, write the reactants on the board, and identify the possible products.
- After the precipitates have settled, record the solunie of precipitate in each precisated cylinder.
- What mode ratio gave the maximum amount of preceptuar? Explain.

#### Part B. Reaction of Copper(II) Chloride with Sodium Phosphate

- Label seven 100-red, gracia and collectors 1-7.
- Using a clean, 50-mil. graduated cylinder, add the appropriate volures of copper(II) chloride solution to each graduated cylinder as shown in Table 2.
- Use a second 50-nd, graduated cylinder to add the appropriate influence of exchanging phosphate solution to each graduated cylinder as shown in Table 2.

Table 2.

Cylinder	1	2	- 3	4.	5	- 60	7
CuCl <sub>2</sub> 0.08 M, mL	10	20	25	340	.3.5	40	540
Na <sub>2</sub> PO <sub>-1</sub> 0.05 M, mL	50	40	35	.340	2.5	1.0	1.0
Cur :PO / Mole Racio	1:5	1:2	2:3	1:1:	3:2	2:1	5:1

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