# 13.2

# THE NATURE OF LIQUIDS

# Section Review

#### Objectives

- Identify factors that determine physical properties of a liquid.
- Define avaporation in terms of kinetic energy
- Describe the equilibrium between a liquid and its vapor
- Identify the conditions under which boiling occurs

#### Vocabulary

- vaporization
- evaporation.
- vapor pressure

- boiling point
- normal boiling point

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Liquids are much \_\_\_\_\_\_ than gases. Liquids and solids

are known as \_\_\_\_\_ states of matter. The conversion of a liquid

to a gas or vapor is called \_\_\_\_\_\_ When a liquid that is not
 \_\_\_\_\_\_ thanges to a gas, the process is called evaporation. A

Liquid evaporates faster when heated; however, evaporation itself

is a \_\_\_\_\_\_ process. When a partially filled container of liquid is

sealed, some of the particles at the \_\_\_\_\_\_ of the liquid vaporize.

These particles collide with the walls of the container, producing

a force called \_\_\_\_\_ The vapor pressure of a liquid can be

determined by a device called a \_\_\_\_\_ A liquid boils when its

\_\_\_\_\_\_ equals the external pressure. The normal boiling point of a

#### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

 Particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize.

# **Nature Of Liquids Section Review Answers**

André van der Braak

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views and ideas expressed at this session and the heated discussion around controversial issues and open questions should lead to the following conclusion Water Science will continue to generate new material of utmost importance and high scientific quality for many more ISOPOW meetings to come

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