Section 5.2 Review The Structure of the Atom

Multiple Choice

For each auestion below, select the letter of the Protect appropriate.

- 1. Which of the following statements about protons, neutrons, and electrons is traw?
 - (a) Protons and electrons have about the same
 - (b) Neutrons and electrons have equal but opposite charges.
 - (c) The mass of a proton is nearly 2000 times greater than the mass of an electron.
 - (d) The mass of an proton is nearly 2000 times greater than the mass of a neutron.
 - (e) Neutrons and electrons are found in the nucleus of an atom.
- 2. What is the atomic number of mercury, Hg?
 - 643 24 ·
 - dbn 800
 - (c) 120
 - (d) 200
 - (c) 200.6
- 3. What is the mass number of an atom of chlorine with 18 neutrons?
 - (a) I-
 - (b) 17.
 - 6cb 18
 - (d) 35
 - (c) 35.5
- 4. What element is represented by the Bohr-Rutherford diagram below?
 - (a) potassium, K.
 - (b) enleium, Ca-
 - (c) hydrogen, H
 - (d) argon, Ar.
 - (c) yttrium, Y

- 5. Which of the following statements about the massesof attoms is folse?
 - (a) Most of the mass of an atom is located in the
 - (b) Electrons contribute very little to the mass of ain attorn.
 - (c) Atoms of different elements have different PERSONAL PROPERTY.
 - (d) Most of the mass of an atom's nucleus is due to the mass of neutrons.
 - (e) The mass number of an atom is equal to the number of protons plus the number of neutrons.
- 6. A neutral atom has eight electrons and eight neutrons. Which of the following options correctly represents the atom?
 - (a) carbon-16
 - (b) oxygen-16
 - (c) oxygen-24
 - (d) oxygen-8
 - (c) sulfur-8
- 7. Which of the following symbols represents the element with an atomic number of 147
 - (a) C
 - Obb IN
 - felt Nil
 - (40) Si
 - tien S.

the statements is false? heutrons an atom has defines

- number delectrons an atom has defines
- The what element it is.

 (d) First the above are true.

 Cone of the above are true. ber of protons an atom has defines

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Jiyuan Zhang

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Table of Contents Mcgraw Hill Ryerson Science 9 Workbook Answers

- 1. Understanding the eBook Mcgraw Hill Ryerson Science 9 Workbook Answers
 - The Rise of Digital Reading Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Ryerson Science 9 Workbook Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Personalized Recommendations
 - Mcgraw Hill Ryerson Science 9 Workbook Answers User Reviews and Ratings
 - Mcgraw Hill Ryerson Science 9 Workbook Answers and Bestseller Lists
- 5. Accessing Mcgraw Hill Ryerson Science 9 Workbook Answers Free and Paid eBooks
 - Mcgraw Hill Ryerson Science 9 Workbook Answers Public Domain eBooks
 - Mcgraw Hill Ryerson Science 9 Workbook Answers eBook Subscription Services
 - Mcgraw Hill Ryerson Science 9 Workbook Answers Budget-Friendly Options

- 6. Navigating Mcgraw Hill Ryerson Science 9 Workbook Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Mcgraw Hill Ryerson Science 9 Workbook Answers Compatibility with Devices
 - Mcgraw Hill Ryerson Science 9 Workbook Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Highlighting and Note-Taking Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Interactive Elements Mcgraw Hill Ryerson Science 9 Workbook Answers
- 8. Staying Engaged with Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Ryerson Science 9 Workbook Answers
- 9. Balancing eBooks and Physical Books Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Mcgraw Hill Ryerson Science 9 Workbook Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Setting Reading Goals Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Ryerson Science 9 Workbook Answers
 - Fact-Checking eBook Content of Mcgraw Hill Ryerson Science 9 Workbook Answers
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- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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