

1.	(a)	(i)	electrons ...16.... protons16 ... neutrons ...16.... All correct (1)	1	
		(ii)	they are isotopes (1) The word 'isotope' or some derivative is essential	1	
		(iii)	S^{2-} / sulphide(2-) ion / sulphur -2 ion / sulphur with two extra electrons (1) but not just S or sulphur	1	
	(b)	(i)	$X(g) + e^- \rightarrow X^-(g)$ ($S(g) + e^- \rightarrow S^-(g)$) (1) $X(g) + e^- \rightarrow X^{2-}(g)$ ($S(g) + e^- \rightarrow S^{2-}(g)$) (1) can have Z in place of X ²⁻ but not Z ²⁻ Penalize omission of state symbol (g) only once.	2	
		(ii)	adding a negative onto a <u>negative species</u> requires energy to overcome repulsion (1) or second electron repelled by negative charge Not second electron repelled by first	1	
		(c)	(Bdecay from) Y <u>produces</u> only one isotope(^{39}Cl) (1) natural sample contains two isotopes (35 and 37) (1)	2	
	2.	(a)	(i)	magnesium / lead / iron (1)	1
			(ii)	iron (1)	1
			(iii)	carbon / phosphorus (1)	1
			(iv)	magnesium / lead (1)	1
(v)			carbon / lead (1)	1	
(b)		Helium (1) because it is the smallest atom (1) with an unshielded nucleus (1) Hydrogen can score max 1 for 'unshielded nucleus' If candidate chooses Carbon, i.e. uses list from part (a), can score 1 mark for reference to 'smallest atom'			3
		(c)	(i)	ability of a positive ion / cation to attract electrons / distort / pull the electron cloud (1) around an anion (1)	2
		(ii)	size (1) Not atomic radius charge (1) charge density worth (2) volume of ion can score in place of size	2	
		(iii)	magnesium chloride magnesium ion smaller than calcium ion (1) polarizes / distorts / attracts electrons around chloride more (1)	2	

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Table of Contents June 2013 Chemistry Unit 5 Mark Scheme

1. Understanding the eBook June 2013 Chemistry Unit 5 Mark Scheme
 - The Rise of Digital Reading June 2013 Chemistry Unit 5 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Chemistry Unit 5 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Chemistry Unit 5 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Chemistry Unit 5 Mark Scheme
 - Personalized Recommendations
 - June 2013 Chemistry Unit 5 Mark Scheme User Reviews and Ratings
 - June 2013 Chemistry Unit 5 Mark Scheme and Bestseller Lists
5. Accessing June 2013 Chemistry Unit 5 Mark Scheme Free and Paid eBooks
 - June 2013 Chemistry Unit 5 Mark Scheme Public Domain eBooks

- June 2013 Chemistry Unit 5 Mark Scheme eBook Subscription Services
- June 2013 Chemistry Unit 5 Mark Scheme Budget-Friendly Options
- 6. Navigating June 2013 Chemistry Unit 5 Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2013 Chemistry Unit 5 Mark Scheme Compatibility with Devices
 - June 2013 Chemistry Unit 5 Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 Chemistry Unit 5 Mark Scheme
 - Highlighting and Note-Taking June 2013 Chemistry Unit 5 Mark Scheme
 - Interactive Elements June 2013 Chemistry Unit 5 Mark Scheme
- 8. Staying Engaged with June 2013 Chemistry Unit 5 Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Chemistry Unit 5 Mark Scheme
- 9. Balancing eBooks and Physical Books June 2013 Chemistry Unit 5 Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Chemistry Unit 5 Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 Chemistry Unit 5 Mark Scheme
 - Setting Reading Goals June 2013 Chemistry Unit 5 Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 Chemistry Unit 5 Mark Scheme
 - Fact-Checking eBook Content of June 2013 Chemistry Unit 5 Mark Scheme
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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