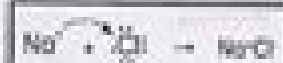


IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to obtain a stable octet of electrons at least part of the time. For example:



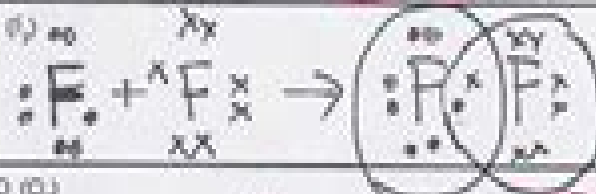
Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



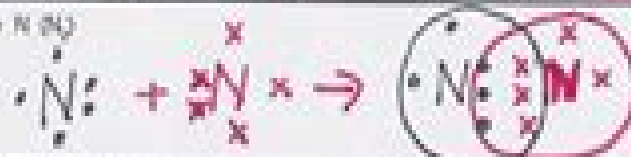
2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO})$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



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Table of Contents Guided Practice Chemistry Chapter 8 Covalent Bonding

1. Understanding the eBook Guided Practice Chemistry Chapter 8 Covalent Bonding
 - The Rise of Digital Reading Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Chemistry Chapter 8 Covalent Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Personalized Recommendations
 - Guided Practice Chemistry Chapter 8 Covalent Bonding User Reviews and Ratings
 - Guided Practice Chemistry Chapter 8 Covalent Bonding and Bestseller Lists
5. Accessing Guided Practice Chemistry Chapter 8 Covalent Bonding Free and Paid eBooks
 - Guided Practice Chemistry Chapter 8 Covalent Bonding Public Domain eBooks
 - Guided Practice Chemistry Chapter 8 Covalent Bonding eBook Subscription Services
 - Guided Practice Chemistry Chapter 8 Covalent Bonding Budget-Friendly Options
6. Navigating Guided Practice Chemistry Chapter 8 Covalent Bonding eBook Formats

- ePub, PDF, MOBI, and More
 - Guided Practice Chemistry Chapter 8 Covalent Bonding Compatibility with Devices
 - Guided Practice Chemistry Chapter 8 Covalent Bonding Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Highlighting and Note-Taking Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Interactive Elements Guided Practice Chemistry Chapter 8 Covalent Bonding
 8. Staying Engaged with Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Chemistry Chapter 8 Covalent Bonding
 9. Balancing eBooks and Physical Books Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Chemistry Chapter 8 Covalent Bonding
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Setting Reading Goals Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Fact-Checking eBook Content of Guided Practice Chemistry Chapter 8 Covalent Bonding
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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