

PHYSICAL SCIENCES - GRADE 10

1 Revision of matter (Grade 9)

1.1 Matter and classification

- Matter is made up of particles (atoms or molecules) and it is the properties of the atoms or molecules that determine the characteristics and reactivity of that matter.
- The properties of matter include: strength; density; melting and boiling points; whether the material is brittle, malleable or ductile; whether it is magnetic or not; and its electrical and thermal conductivity.
- Elements and compounds are classified as pure because they contain particles that are all the same. Substances can be classified as pure – elements and compounds
 - or mixtures. Pure substances contain particles that are all the same. Elements and compounds can be represented by symbols and formulae. Some of the names, symbols or formulae of a number of common elements and compounds that you should have learnt are the following:
 - Elements: S for sulphur; C for carbon; P for phosphorus (note the spelling); H for hydrogen; O for oxygen; He for helium; N for nitrogen; F for fluorine; Mg for magnesium; Ca for calcium; Zn for zinc.
 - Compounds: H_2O for water; H_2SO_4 for sulphuric acid; HCl for hydrochloric acid; NaCl for sodium chloride (table salt); $CaCO_3$ for calcium carbonate; KNO_3 for potassium nitrate; Na_3PO_4 for sodium phosphate.
- You will notice that the naming of compounds follows these rules
 - The metal or hydrogen is always named first, then the non-metal (potassium iodide: KI; hydrogen chloride: HCl (also called hydrochloric acid)).
 - The name of a non-metal ion of an element always ends in -ide (for instance, sodium chloride: NaCl).
 - The name of a non-metal complex ion (more than one element in the ion) usually ends in -ate or -ite (for instance, copper sulphate: $CuSO_4$; potassium sulphite: K_2SO_3). An exception to this is hydroxide (for instance, sodium hydroxide: NaOH).
- Mixtures consist of different substances, and can be homogeneous (have the same 'look' throughout) or heterogeneous (you can see the different substances). They can be separated in various ways, for instance filtering, sieving, chromatography, distilling or dissolving one constituent.
- Some of the properties of metals are that they conduct electricity and heat, and they are malleable and ductile.
- Non-metals are non-conductors of electricity and heat. They are gases, liquids or brittle solids, and bond with each other.
- Metalloids are elements on the border between metals and non-metals, and they have both metal and non-metal properties.

Guided Outline Ch 1physical Science

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Guided Outline Ch 1physical Science:

The book delves into Guided Outline Ch 1physical Science. Guided Outline Ch 1physical Science is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Guided Outline Ch 1physical Science, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Guided Outline Ch 1physical Science
 - Chapter 2: Essential Elements of Guided Outline Ch 1physical Science
 - Chapter 3: Guided Outline Ch 1physical Science in Everyday Life
 - Chapter 4: Guided Outline Ch 1physical Science in Specific Contexts
 - Chapter 5: Conclusion
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6. In chapter 5, this book will draw a conclusion about Guided Outline Ch 1physical Science. This chapter will summarize the key points that have been discussed throughout the book.

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Table of Contents Guided Outline Ch 1physical Science

1. Understanding the eBook Guided Outline Ch 1physical Science
 - The Rise of Digital Reading Guided Outline Ch 1physical Science
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Outline Ch 1physical Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Outline Ch 1physical Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Outline Ch 1physical Science
 - Personalized Recommendations
 - Guided Outline Ch 1physical Science User Reviews and Ratings
 - Guided Outline Ch 1physical Science and Bestseller Lists
5. Accessing Guided Outline Ch 1physical Science Free and Paid eBooks
 - Guided Outline Ch 1physical Science Public Domain eBooks
 - Guided Outline Ch 1physical Science eBook Subscription Services
 - Guided Outline Ch 1physical Science Budget-Friendly Options
6. Navigating Guided Outline Ch 1physical Science eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Outline Ch 1physical Science Compatibility with Devices
 - Guided Outline Ch 1physical Science Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Outline Ch 1physical Science
 - Highlighting and Note-Taking Guided Outline Ch 1physical Science
 - Interactive Elements Guided Outline Ch 1physical Science

8. Staying Engaged with Guided Outline Ch 1physical Science
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Outline Ch 1physical Science
9. Balancing eBooks and Physical Books Guided Outline Ch 1physical Science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Outline Ch 1physical Science
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Outline Ch 1physical Science
 - Setting Reading Goals Guided Outline Ch 1physical Science
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Outline Ch 1physical Science
 - Fact-Checking eBook Content of Guided Outline Ch 1physical Science
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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