

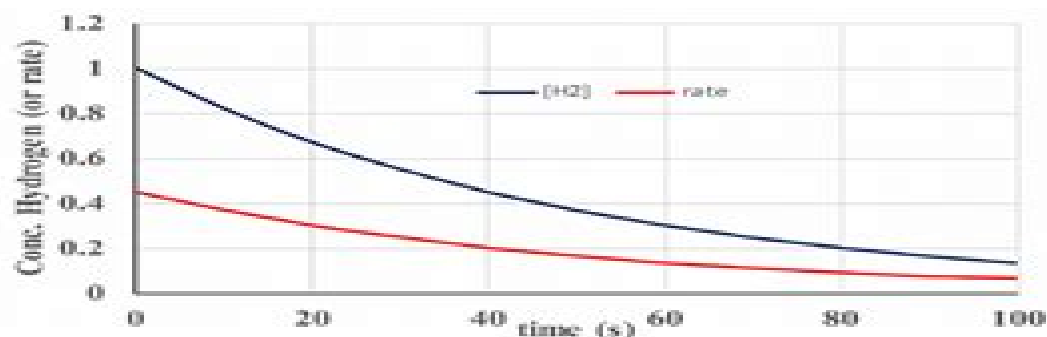
## Chapter 15

### Chemical Kinetics

1. The rate of chemical reaction – the change in the amounts of reagents (or products) divided by the change in time.

For the reaction of hydrogen and iodine,  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightarrow 2 \text{HI}(\text{g})$ , the reaction rate is defined as

Rate =  $-\frac{\Delta[\text{H}_2]}{\Delta t} = -\frac{\Delta[\text{I}_2]}{\Delta t} = +\frac{1}{2} \frac{\Delta[\text{HI}]}{\Delta t}$  Note that reaction rate is a positive number, therefore, negative sign must be added when reagents are used (because their concentrations decrease) and positive sign when products are used (because concentrations of products increase). Factor  $\frac{1}{2}$  is due to stoichiometry because two moles of HI are made from one mole of  $\text{H}_2$  and  $\text{I}_2$ . Concentration of hydrogen and the reaction rate (magnified 25 times) for the above reaction, is shown below



2. The instantaneous reaction rate - this is the slope of the concentration curve plotted against the reaction time. For general reaction:  $a\text{A} + b\text{B} \rightarrow c\text{C} + d\text{D}$ , the reaction rate is defined as:  
Rate =  $-\frac{1}{a} \frac{\Delta[\text{A}]}{\Delta t} = -\frac{1}{b} \frac{\Delta[\text{B}]}{\Delta t} = +\frac{1}{c} \frac{\Delta[\text{C}]}{\Delta t} = +\frac{1}{d} \frac{\Delta[\text{D}]}{\Delta t}$
3. Measuring reaction rates  $\rightarrow$  determine concentration of reagents or products as a function of time, using suitable analytical techniques, for example, absorption of light or chromatography.
4. The rate laws: The reaction rate depends on concentration of reagents. Let's consider reaction  $\text{A} \rightarrow \text{products}$  (for example, decomposition of sucrose into glucose and

# General Chemistry Ii Chapter Chemical Kinetics Answers

**Daniela Niemeyer**



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to specific sites chemotaxis catalytically induced self propulsion polymerization growth of dendrites or aggregates or nuclei of a new phase Several decades ago it was recognized that the kinetic behavior in some systems with reactions and random transport is strongly affected by many factors which were not taken into account in previous studies These are to name but a few fluctuations in the spatial distributions of the reactants and fluctuations of the reactivity some essentially many particle phenomena effects of anomalous diffusion molecular crowding as well as the internal geometry of the reaction bath Within recent years along with a growing interest in chemical processes occurring in biological systems or cellular environments numerous advances have been made and considerable knowledge has been acquired These seminal contributions are however scattered among many journals and no attempt has been made so far to present a unified picture This book presents a general overview of different contemporary facets of chemical kinetics in a variety of different environments It includes 23 seminal works and reviews on different aspects of reaction processes in chemical physical and biophysical systems both theoretical and experimental

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