F321: Atoms, Bonds and Groups Redox

39 Marks

 The element strontium forms a nitrate, Sr(NO₃)₂, which decomposes on heating as shown below.

 $2Sr(NO_3)_2(s) \rightarrow 2SrO(s) + 4NO_2(g) + O_2(g)$

- (i) Using exidation numbers, explain why the reaction involves both exidation and reduction.
- (ii) A student heats 5.29 g of Sr(NO₃)₂ and collects the gas at room temperature and pressure, RTP.

$$2Sr(NO_3)_2(s) \rightarrow 2SrO(s) + 4NO_2(g) + O_2(g)$$

Calculate the volume of gas, in \mbox{dm}^3 , obtained by the student at RTP.

Molar mass of $Sr(NO_3)_2 = 211.6 g mol^{-1}$.

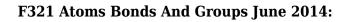
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[Total 6 marks]

121

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