

# Experiment 8 Report Sheet

## Limiting Reactant

Date \_\_\_\_\_

### A. Precipitation of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ from the Salt Mixture

Unknown number B

1. Mass of beaker (g)
2. Mass of beaker and salt mixture (g)
3. Mass of salt mixture (g)
4. Mass of filter paper (g)
5. Mass of filter paper and product after air-dried or oven-dried (g)
6. Mass of dried product (g) 1.191
7. Formula of dried product

### B. Determination of Limiting Reactant

1. Limiting reactant in salt mixture (write complete formula)
2. Excess reactant in salt mixture (write complete formula)

### Data Analysis

1. Moles of  $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$  (or  $\text{CaC}_2\text{O}_4$ ) precipitated (mol)
2. Moles of limiting reactant in salt mixture (mol)  
• formula of limiting hydrate  $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$
3. Mass of limiting reactant in salt mixture (g)
4. Mass of excess reactant in salt mixture (g)  
• formula of excess hydrate  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
5. Percent limiting reactant in salt mixture (%)
6. Percent excess reactant in salt mixture (%)
7. Mass of excess reactant that reacted (g)
8. Mass of excess reactant, unreacted (g)

\*Show calculations for Trial 1 on next page.

Trial 1 250 mL

Trial 2

96.553g

47.956g

97.559g

49.083g

1.001g

1.067g

0.785g

0.785g

1.976g

1.774g

1.191g

0.989g

$\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$

$\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$

$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$

$\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$

0.00815 mol

0.00676 mol

2. Moles of limiting reactant in salt mixture (mol)

• formula of limiting hydrate  $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$

0.00815 mol

0.00676 mol

3. Mass of limiting reactant in salt mixture (g)

1.198g  $\text{CaCl}_2$

0.994g  $\text{CaCl}_2$

4. Mass of excess reactant in salt mixture (g)

• formula of excess hydrate  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$

0.192g

0.073g

5. Percent limiting reactant in salt mixture (%)

93.97%

93.16%

6. Percent excess reactant in salt mixture (%)

19.32%

6.84%

7. Mass of excess reactant that reacted (g)

1.4922g

1.2343g

8. Mass of excess reactant, unreacted (g)

1.3030g

1.1606g

47.02g/  
1 mole

7.02g  
x  
0.00815

$$\textcircled{D} \quad \frac{1.191g \text{ CaC}_2\text{O}_4}{146.12g} \times 1 \text{ mol}$$

$$\frac{0.989g}{146.12g} \times 1 \text{ mol}$$

# Experiment 8 Limiting Reactant Answers

**Thirumalaisamy P.**  
**Velavancorresponding**



## **Experiment 8 Limiting Reactant Answers:**

## Embracing the Melody of Appearance: An Mental Symphony within **Experiment 8 Limiting Reactant Answers**

In a world eaten by screens and the ceaseless chatter of immediate conversation, the melodic beauty and emotional symphony created by the published term frequently disappear in to the background, eclipsed by the constant sound and disruptions that permeate our lives. However, located within the pages of **Experiment 8 Limiting Reactant Answers** a marvelous fictional treasure brimming with natural thoughts, lies an immersive symphony waiting to be embraced. Constructed by a masterful musician of language, this fascinating masterpiece conducts viewers on an emotional trip, well unraveling the concealed tunes and profound influence resonating within each carefully crafted phrase. Within the depths of the moving evaluation, we can examine the book is main harmonies, analyze its enthralling publishing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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