

**Experiment 23****Advance Study Assignment:** Determination of the Equilibrium Constant for a Chemical Reaction

1. A student mixes 5.00 mL  $2.00 \times 10^{-3}$  M  $\text{Fe}(\text{NO}_3)_3$  in 1 M  $\text{HNO}_3$  with 3.00 mL  $2.00 \times 10^{-3}$  M  $\text{KSCN}$  and 2.00 mL of water. She finds that in the equilibrium mixture the concentration of  $\text{FeSCN}^{2+}$  is  $7.0 \times 10^{-5}$  M. Find  $K_c$  for the reaction  $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$ .

**Step 1** Find the number of moles  $\text{Fe}^{3+}$  and  $\text{SCN}^{-}$  initially present. (Use Eq. 3.)

$V_{\text{Fe}^{3+}}$

\_\_\_\_\_ moles  $\text{Fe}^{3+}$ ; \_\_\_\_\_ moles  $\text{SCN}^{-}$

**Step 2** How many moles of  $\text{FeSCN}^{2+}$  are in the mixture at equilibrium? What is the volume of the equilibrium mixture? (Use Eq. 3.)

\_\_\_\_\_ mL; \_\_\_\_\_ moles  $\text{FeSCN}^{2+}$

How many moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^{-}$  are used up in making the  $\text{FeSCN}^{2+}$ ?

\_\_\_\_\_ moles  $\text{Fe}^{3+}$ ; \_\_\_\_\_ moles  $\text{SCN}^{-}$

**Step 3** How many moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^{-}$  remain in the solution at equilibrium? (Use Eq. 4 and the results of Steps 1 and 2.)

\_\_\_\_\_ moles  $\text{Fe}^{3+}$ ; \_\_\_\_\_ moles  $\text{SCN}^{-}$

**Step 4** What are the concentrations of  $\text{Fe}^{3+}$ ,  $\text{SCN}^{-}$ , and  $\text{FeSCN}^{2+}$  at equilibrium? (Use Eq. 3 and the results of Step 2 and Step 3.)

$[\text{Fe}^{3+}] =$  \_\_\_\_\_ M;  $[\text{SCN}^{-}] =$  \_\_\_\_\_ M;  $[\text{FeSCN}^{2+}] =$  \_\_\_\_\_ M

**Step 5** What is the value of  $K_c$  for the reaction? (Use Eq. 2 and the results of Step 4.)

$K_c =$  \_\_\_\_\_

(continued on following page)

# Experiment 23 Determination Equilibrium Constant Answers

**ML Morrison**



## **Experiment 23 Determination Equilibrium Constant Answers:**

The Enigmatic Realm of **Experiment 23 Determination Equilibrium Constant Answers**: Unleashing the Language is Inner Magic

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