

$$\Delta S [\text{H}_2\text{O}_2 (g)] = 232.9 \text{ J K}^{-1} \text{ mol}^{-1} \quad \Delta H_f [\text{H}_2\text{O}_2 (g)] = -136.1 \text{ kJ mol}^{-1}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}} = 5.3 \text{ J K}^{-1} \text{ mol}^{-1} \text{ spontaneous}$$

7. The reaction $\text{NH}_3(g) + \text{HCl}(g) \rightarrow \text{NH}_4\text{Cl}(s)$ is spontaneous at room temperature. Determine the temperature at which it is no longer spontaneous. Given information

$$S^\circ [\text{NH}_3(g)] = 193.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{HCl}(g)] = 187.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{NH}_4\text{Cl}(s)] = 94.56 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_3(g)] = -46.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{HCl}(g)] = -92.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_4\text{Cl}(s)] = -315.39 \text{ kJ mol}^{-1}$$

$$T < 619 \text{ K}$$

8. Calculate the change in the entropies of the system and the surroundings, and the total change in entropy, when the volume of a sample of argon gas of mass 21 g at 298 K and 1.50 bar increases from 1.20 dm³ to 4.60 dm³ in (a) an isothermal reversible expansion, (b) an isothermal irreversible expansion against $p_{\text{ex}} = 0$, and (c) an adiabatic reversible expansion

$$\text{a) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = 0 \text{ J/K}$$

$$\text{b) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +3.0 \text{ J/K}$$

$$\text{c) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +3.0 \text{ J/K}$$

9. The standard molar enthalpy of formation of $\text{H}_2\text{O}(g)$ is $-241.8 \text{ kJ mol}^{-1}$ and its molar entropy is $188.8 \text{ J K}^{-1} \text{ mol}^{-1}$ at 298 K. Calculate the standard molar enthalpy of formation of $\text{H}_2\text{O}(l)$ at 298 K.

Standard enthalpy of formation of $\text{H}_2\text{O}(l)$ is $-285.8 \text{ kJ mol}^{-1}$. The standard molar enthalpy of vaporization of H_2O is 44.0 kJ mol^{-1} at 298 K.

$$\Delta H_f^\circ [\text{H}_2\text{O}(l)] = -285.8 \text{ kJ mol}^{-1}$$

$$\Delta H_{\text{vap}}^\circ [\text{H}_2\text{O}(l)] = 44.0 \text{ kJ mol}^{-1}$$

10. The standard molar enthalpy of formation of $\text{H}_2\text{O}(g)$ is $-241.8 \text{ kJ mol}^{-1}$ and its molar entropy is $188.8 \text{ J K}^{-1} \text{ mol}^{-1}$ at 298 K. The standard molar enthalpy of formation of $\text{H}_2\text{O}(l)$ is $-285.8 \text{ kJ mol}^{-1}$ at 298 K. Calculate the standard molar enthalpy of formation of $\text{H}_2\text{O}(s)$ at 298 K.

$$\Delta H_f^\circ [\text{H}_2\text{O}(s)] = -291.8 \text{ kJ mol}^{-1}$$

11. Suppose you had 1.00 mol of $\text{H}_2\text{O}(g)$ at 298 K and 1.00 bar. Calculate the standard molar enthalpy of formation of $\text{H}_2\text{O}(l)$ at 298 K.

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Y. V. C. Rao



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