

Directions: Read carefully and solve the following problems.

1. The complete combustion of Octane results to the production of Carbon dioxide and water as shown in the unbalanced chemical equation below.



- Write the balanced chemical reaction equation.
 - How much CO_2 in grams will be produced from burning 30 grams of Octane?
 - What mass of Oxygen is needed to completely consume 100 grams of Octane?
2. In a certain experiment, Hydrochloric acid is used to decompose Calcium carbonate. The chemical reaction equation is shown below.



- After the experiment, the total mass of CaCl_2 produced is 120g. Assuming that the reaction proceeded to completion, how many moles of CaCO_3 were used in the experiment?
 - Based on the amount of CaCl_2 produced in the reaction, how much CO_2 in grams effervesced out of the Calcium carbonate used?
3. Hydrogen fluoride is produced by reacting Calcium fluoride with sulfuric acid. A by-product of the process is Calcium sulfate.

Exercise 11 Stoichiometry

**Morris Hein, Leo R. Best, Robert L.
Miner**



Exercise 11 Stoichiometry:

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Solid-State Electrochemistry Abdelkader Hammou, Samuel Georges, 2020-02-03 This book features the essential material for any graduate or advanced undergraduate course covering solid state electrochemistry It provides the reader with fundamental course notes and numerous solved exercises making it an invaluable guide and compendium for students of the subject The book places particular emphasis on enhancing the reader s expertise and comprehension of thermodynamics the Kr ger Vink notation the variation in stoichiometry in ionic compounds and of the different types of electrochemical measurements together with their technological applications Containing almost 100 illustrations a glossary and a bibliography the book is particularly useful for Master and PhD students industry engineers university instructors and researchers working with inorganic solids in general

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Chemistry: Core Concepts, 3rd Edition Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package

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Resources in Education, 1980

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