Chapter 11 (continued)

$$\begin{array}{c} 2 \, H_2 O_2(l) \to O_2(g) + 2 \, H_2 O(l) \\ \frac{2 \, \operatorname{mol} \, H_2 O_2}{1 \, \operatorname{mol} \, O_2} \quad \text{and} \quad \frac{2 \, \operatorname{mol} \, H_2 O_2}{2 \, \operatorname{mol} \, H_2 O} \\ \\ \frac{1 \, \operatorname{mol} \, O_2}{2 \, \operatorname{mol} \, H_2 O_2} \quad \text{and} \quad \frac{1 \, \operatorname{mol} \, O_2}{2 \, \operatorname{mol} \, H_2 O} \\ \\ \frac{2 \, \operatorname{mol} \, H_2 O_2}{2 \, \operatorname{mol} \, H_2 O_2} \quad \text{and} \quad \frac{2 \, \operatorname{mol} \, H_2 O}{1 \, \operatorname{mol} \, O_2} \end{array}$$

To determine the number of mole ratios that defines a given chemical reaction, multiply the number of species in the equation by the next lower number. Thus, a chemical reaction with three participating species can be defined by six mole ratios ($3 \times 2 = 6$); a chemical reaction with four species can be defined by 12 mole ratios ($4 \times 3 = 12$).

Why learn to write mole ratios? They are the key to calculations that are based on chemical equations. Using a balanced chemical equation, mole ratios derived from the equation, and a given amount of one of the reactants or products, you can calculate the amount of any other participant in the reaction.

Practice Problems

- Determine all the mole ratios for the following balanced chemical equations.
 - a. $N_2(g) + O_2(g) \rightarrow 2NO(g)$
 - **b.** $4NH_3(aq) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(1)$
 - c. $4HCl(aq) + O_2(g) \rightarrow 2H_2O(l) + 2Cl_2(g)$

11.2 Stoichiometric Calculations

There are three basic stoichiometric calculations: mole-to-mole conversions, mole-to-mass conversions, and mass-to-mass conversions. All stoichiometric calculations begin with a balanced equation and mole ratios.

► Stoichiometric mole-to-mole conversion How can you determine the number of moles of table salt (NaCl) produced from 0.02 moles of chlorine (Cl₂)?

Exercise 11 Stoichiometry 1

Lei Shi

Exercise 11 Stoichiometry 1:

Whispering the Strategies of Language: An Psychological Quest through Exercise 11 Stoichiometry 1

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