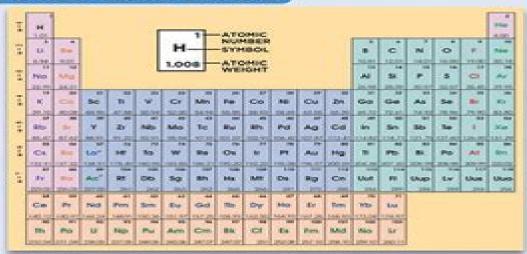




Periodic Table of Elements



as If the property of the property

- Monatomic: Noble gueses: belians, noon, argon, krypton, senon, and radion; monteactive
- Distance: General fluorine, chilorine, hydrogen, oxygen, and nitrogen; Liquid: becoming: Solid: inding
- Mutecular solids: Sulfor
- Considerat unlider Carbon (diamond).
- Metals: Most common type of element; high coordination numbers; worst democ
- Types of compounds: Lurge variety: due to lonic, covalent, and metallic bonding options.
 - Bomier, Elloctrom exchange.
 - Covalent: Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex: Central ion # Solvent. Transition metal compounds:
 - Coordination compound: Non-
 - bonding Strands Organometallie: Metal-C bond
- Complex minerals: Metal exides mental malling attaches
- · Periodicity: The chemical properties of an element depend on valence electrons.
 - Family of elements: Column in. periodic table.
 - The addition of d-orbitals produces. am expranded vertet.

Periodic Trends

- · Chemistry determined by the behavior of valence electrons.
- Atoms or ions got larger as you go down the column; electron shells are added
- Cation and anions (same charec) decrease in size as the atomic number (Z) increases.

- Ionization energies: Noble gases are: the most stable; alkali metals are the easiest to lonize; electron screening lessons Z-effect on orbital energy
- Atomic and looke radii: Not real properties; derive from structural data; divide bond into radii using a consistent and Properties
 - · Ionic rudius: Depends on charge, counter ion, and coordination terratail types).
 - Attomic radius: Based on diatomic gas data or metallic solid structural distant

Development of the Periodic Table

- Missing elements were prodicted using: periodicity
- Culumns In Alkali metals.
- Column 2: Alkaline earth metals
- · First transition series: Titatium-
- Column 17: Welegons: flooring, chiorine, bromine, indine, and autatine
- Culumn 18: Rary penese belium, neon, argon, keypton, senon, and radios
- Actinides: Z = 90 or above; radioactive Lanthanides: f-subshell elements:



Practical Application















Atomic Structure

- Atomic number (Z): The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge
- -q, Z+q electrons; for a cution of charge +q, Z-q electrons. Mass number (A): A=Z+N, where N = the number of neutrons in the mucleus.
- Instance: Atoms with the same Z. different A: an element can have: one or more instances, any nample contains a number of instances; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)





Atomic Quantum Numbers & Orbitals

ms principle.

& angular momentum (orbital shape)

max magnetic (orbital direction) max electron spin-■ For each n, the possible / values are 0, 1,..., n - 1

For each L the possible m, values are -L ... 0, +1

m, has two possible values: #15 and -15 (spin up, spin down). · Each value of a denotes a "shell" in the atomic structure; each

Edicators a subshelli. a-carbitals I = 0, 1 type d-orthitals: I = 2, 5 types

p-orbitals I = 1, 3 types fi-orthitals t = 3, 2 types

Many-Electron Atoms

Orbinals and quantum num bers derived for the hydrogen atom are used to describe all munty-electron atoms and konse a more rigorous treatment determines exact everyslevels and orbital proporties; the main result is that subshells have different energies; the filling of levels is guided by the Authors principle:



•		-	-	- Allendaria		
1	- 0	1000	10.00	31	I s-orbital	
1	- 0	- 0	-16	3 Sec.	aphorical	
2	0	- 0	950	1.0	2x-orbital	
2	0	- 0	-85	Die	sphorical	
2		1	456	410	2p-orbitals	
	1	- 0	-56	C		
2	1	+1	416	SIN	Introdut	
2	1	-10	3486	-0	W. D. D.	
20	1000	0.7	-35	100	200000000000000000000000000000000000000	
2:	1	+1	-86	C 396at -		

- Electron configuration: Atomic orbital occupancy for an atom.
- Pauli exclusion principle: Each electron has a unique set of quantum numbers; as orbital may hold up to two electrons; one with spin up, m. = +15, and one with spin down, m = -15
- Hand's rule: Electrons fill p. d. and f subshells in a manner to maximize the number of unpoined electrons thalf-filled orbitals); medica but maken head made
- Indication potential (IP): Energy required to remove electron finden are above or lices.
 - · First IP: Removal of first valence electron
- Electromegativity: Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

Lavc Chemistry Study Guide

F Rizvi

Lavc Chemistry Study Guide:

Decoding Lavc Chemistry Study Guide: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Lavc Chemistry Study Guide**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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