

Section A

For each question there are four possible answers, A, B, C and D. Choose the one you consider to be correct.

- 1 Which equation correctly represents the enthalpy change of atomisation of iodine?



- 2 What is the second electron affinity of oxygen?
Use the data below to determine your answer.

| | $\Delta H \text{ kJ/mol}$ |
|--|---------------------------|
| Enthalpy of formation of calcium oxide | -636 |
| Lattice enthalpy of calcium oxide | -3401 |
| Second ionisation energy of calcium | +1100 |
| First ionisation energy of calcium | +590 |
| Enthalpy of atomisation of calcium | +177 |
| Enthalpy of atomisation of oxygen | +249 |
| First electron affinity of oxygen | -141 |

A $+790 \text{ kJ/mol}$

B -790 kJ/mol

C $+1\,039 \text{ kJ/mol}$

D $-1\,039 \text{ kJ/mol}$

- 3 Why does the enthalpy change of neutralisation of oleic acid, $\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{CO}_2\text{H}$, with aqueous ammonia differ from the theoretical value of $-57.2 \text{ kJ/mol}^{-1}$?

A Oleic acid and ammonia do not react in the ratio 1:1.

B Oleic acid is unsaturated therefore its enthalpy of neutralisation is more exothermic.

C Energy is required to ionise oleic acid and ammonia.

D Ammonia is completely hydrolysed.

June 2013 Chemistry Paper 3

Shasha Hu



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