

STANDARD FEATURES—5410 AND 5510

•John Deere 4000 Series Engine

- 5410 CD4045DLV50 56 kw (75 hp)
- 5510 CD4045TLV50 63 kw (85 hp)

4-cylinder diesel engine.

- Wet sleeved.
- Direct injection.
- Intake air heater starting aid.
- Key switch controlled fuel shut-off.
- 5410 is naturally aspirated and the 5510 is turbocharged.

•CollarShift, Transmission

- Nine speeds forward, three reverse
- Inboard planetary final drives
- Differential lock

•Dual Clutch

- Provides continuous live PTO
- Stops tractor without disengaging PTO

•PTO

- Rear, 540 rpm
- 540/540E PTO available on SyncShuttle™ transmission only
- Fully independent clutch

•Hydrostatic Power Steering

- Power is supplied by a tandem gear hydraulic pump mounted to the engine

•Tilt Steering Wheel

•Hydraulic Brakes

- Wet disc
- Individually hydraulic controlled
- Self-adjusting
- Self-equalizing

•Open-Center Hydraulic System

- Tandem gear hydraulic pumps
- Pumps are driven directly off engine timing gears

The rear pump supplies oil to the power steering, provides brake valve make-up oil and lubricates top shaft of the transmission.

The front pump supplies oil to the rockshaft and the selective control valves, if equipped.



LV1888

Side LV1888



LV1890

Side LV1890



Side LV1891

John Deere 5410 Repair Manual

N Noddings



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O $2\text{Cr}^{3+} \text{aq}$ $2\text{OH}^- \text{aq}$

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq}$ $2\text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2\text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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canceled to give the net ionic equation $3\text{Ba}^{2+} + 2\text{PO}_4^{3-} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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$\text{AgNO}_3 + \text{KCl} + \text{AgCl} + \text{KNO}_3$ total ionic $\text{Ag}^+ + \text{NO}_3^- + \text{K}^+ + \text{Cl}^- + \text{AgCl} + \text{K}^+ + \text{NO}_3^-$ net ionic $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} + \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{l}$ the spectator ions in $\text{KOH} + \text{HNO}_3$ are K^+ and NO_3^- how to identify and list the

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