

SOME BASIC CONCEPTS OF PHYSICAL CHEMISTRY

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Measurement in chemical (Precision, significant figures, SI units, Dimensional analysis), Laws of chemical combination, Atomic Mass, Molecular Mass, molar concept, Molar Mass, determination of Molecular formula, Chemical equation, stoichiometry of Chemical reactions.

INITIAL STEP EXERCISE

Multiple Choice Questions

1. 1 g of the carbonate of a metal was dissolved in 25 ml of HCl . The resulting liquid required 8 ml of N-NaOH for neutralisation. The eq. wt. of the metal carbonate is
 (a) 20 (b) 30
 (c) 40 (d) None
2. 8 ml of N-HCl , 20 ml of $\text{Na}_2\text{B}_4\text{O}_7$ and 20 ml of $\text{NaOH} + \text{HNO}_3$ are mixed together and the mixture made to 1 litre.
 (a) The normality of the resulting solution is
 (a) N/5 (b) N/10
 (c) N/20 (d) N/40
 (b) The vol. of pure NaOH required to neutralise the above solution is
 (a) 10 g (b) 5 g
 (c) 1 g (d) 2.5 g
3. If 0.1 mol of BaCl_2 is mixed with 0.20 mol of Na_3PO_4 , the maximum amount of $\text{Ba}_3(\text{PO}_4)_2$ that can be formed is
 (a) 0.10 mol (b) 0.20 mol
 (c) 0.25 mol (d) 0.15 mol
4. The equivalent mass of MnO_2 is half of its molar mass when it is converted to
 (a) Mn_2O_3 (b) MnCl_2
 (c) MnCl_2^{+} (d) MnCl_2^{-}
5. The nitrate nitride can be converted into ammonium ion. The equivalent mass of NO_2^{-} ion in this reaction would be
 (a) 4.20 g (b) 7.20 g
 (c) 10.2 g (d) 22.8 g

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Table of Contents Inorganic Chemistry Multiple Choice Exams

1. Understanding the eBook Inorganic Chemistry Multiple Choice Exams
 - The Rise of Digital Reading Inorganic Chemistry Multiple Choice Exams
 - Advantages of eBooks Over Traditional Books
2. Identifying Inorganic Chemistry Multiple Choice Exams
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Inorganic Chemistry Multiple Choice Exams
 - User-Friendly Interface
4. Exploring eBook Recommendations from Inorganic Chemistry Multiple Choice Exams

- Personalized Recommendations
 - Inorganic Chemistry Multiple Choice Exams User Reviews and Ratings
 - Inorganic Chemistry Multiple Choice Exams and Bestseller Lists
5. Accessing Inorganic Chemistry Multiple Choice Exams Free and Paid eBooks
 - Inorganic Chemistry Multiple Choice Exams Public Domain eBooks
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 6. Navigating Inorganic Chemistry Multiple Choice Exams eBook Formats
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 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Inorganic Chemistry Multiple Choice Exams
 - Highlighting and Note-Taking Inorganic Chemistry Multiple Choice Exams
 - Interactive Elements Inorganic Chemistry Multiple Choice Exams
 8. Staying Engaged with Inorganic Chemistry Multiple Choice Exams
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Inorganic Chemistry Multiple Choice Exams
 9. Balancing eBooks and Physical Books Inorganic Chemistry Multiple Choice Exams
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Inorganic Chemistry Multiple Choice Exams
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Inorganic Chemistry Multiple Choice Exams
 - Setting Reading Goals Inorganic Chemistry Multiple Choice Exams
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Inorganic Chemistry Multiple Choice Exams

- Fact-Checking eBook Content of Inorganic Chemistry Multiple Choice Exams
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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