

সুতরাং, বিশেষ MY_2 অক্সিজেন হবে।



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কিছু ভাষ্যমাত্রায় কোনো প্রকণে একটি লবণ প্রস্তুত থাকবে যদি অক্সিজেন হবে তা নির্ভর করে লবণটির প্রাকৃতিক তপনতের উপর। যদি প্রকণে কোনো লবণের উপস্থিতি অক্সিজেনের আয়নিক তপনতের মান লবণটির প্রাকৃতিক তপনতের মানকে অতিক্রম করতে না পারে তবে লবণটি প্রকণে প্রস্তুত অকৃত্য থাকবে। আর যদি লবণের আয়নিক তপনতের মান (K_{sp}) লবণটির প্রাকৃতিক তপনতের মান (K_{sp}) কে অতিক্রম করে তবে লবণটি প্রকণ থেকে অক্সিজেন হবে।

অর্থাৎ,

$K_{sp} < K_{sp}$: প্রকণটি অসম্পূর্ণ প্রকণ

$K_{sp} = K_{sp}$: প্রকণটি সম্পূর্ণ প্রকণ

$K_{sp} > K_{sp}$: প্রকণ থেকে প্রব অক্সিজেন হবে

III A পরে PN প্রকণ যোগ করলে তা নিম্নলিখিত অক্সিজেন হয়,

PN Hooooo P⁺ + N⁻

প্রকণে, PN প্রকণের ঘনত্ব, 0.1M হবে-

P⁺ আয়নের ঘনত্ব, [P⁺] = 0.1 M

N⁻ আয়নের ঘনত্ব, [N⁻]₁ = 0.1 M

এখানে, PN প্রকণে N⁻ আয়ন উপস্থিত। অন্যদিকে A পরে MN₃ লবণের একটি অক্সিজেন N⁻ উপস্থিত থাকবে A পরে PN প্রকণ যোগ করলে সমস্ত অক্সিজেন প্রকণের করলে MN₃ এর প্রাকৃতিক প্রকণ হবে।

কি, MN₃ এর লবণ প্রাকৃতিক, S

প্রকণে MN₃ নিম্নলিখিত অক্সিজেন হয়,

MN₃ Hooooo M³⁺ + 3N⁻

সুতরাং, M³⁺ আয়নের ঘনত্ব, [M³⁺] = S

N⁻ আয়নের ঘনত্ব, [N⁻]₂ = 3S

সেহেতু, N⁻ আয়ন PN ও MN₃ উভয় দিকে উপস্থিত, সেহেতু

N⁻ আয়নের মোট ঘনত্ব হবে, [N⁻] = (3S + 0.1)M

এখন MN₃ এর আয়নিক তপনত,

$$K_{sp} = [M^{3+}] \times [N^-]^3$$

$$\begin{aligned} \text{কি, } 2.7 \times 10^{-11} &= S \times (3S + 0.1)^3 \quad [\because \text{কি, } K_{sp} = 2.7 \times 10^{-11}] \\ &= S \times [(3S)^3 + 3 \times (3S)^2 \times 0.1 + 3 \times 3S \times (0.1)^2 + (0.1)^3] \\ &= S \times [27S^3 + 2.7S^2 + 0.09S + 0.001] \\ &= 27S^4 + 2.7S^3 + 0.09S^2 + 0.001S \end{aligned}$$

$$\text{কি, } 2.7 \times 10^{-11} = 0.001S \quad [\because S^4, S^3, S^2 \text{ এর মান খুবই নগণ্য}]$$

$$\begin{aligned} \therefore S &= \frac{2.7 \times 10^{-11}}{0.001} \\ &= 2.7 \times 10^{-8} \text{ mol L}^{-1} \end{aligned}$$

অর্থাৎ, MN₃ দিকের লবণ প্রাকৃতিক, $2.7 \times 10^{-8} \text{ mol L}^{-1}$

অতএব, A পরে PN প্রকণ যোগ করলে MN₃ এর প্রাকৃতিক $10^{-8} \text{ mol L}^{-1}$ থেকে PN পরে $2.7 \times 10^{-8} \text{ mol L}^{-1}$ হবে।

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Lingsheng Yao



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