

STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2013

[Marks : 70

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) In body centred cubic structure the space occupied is about _____
 (a) 68% (b) 53% (c) 38% (d) 32%
- (ii) For a gaseous reaction the unit of rate of reaction is _____
 (a) L atm s⁻¹ (b) atm mol⁻¹ s⁻¹ (c) atm s⁻¹ (d) mol s
- (iii) Which of the following compounds contains S = O as well as S = S bonds?
 (a) Sulphuric acid (b) Thiosulphuric acid
 (c) Sulphurous acid (d) Thiosulphurous acid
- (iv) Which of the following solutions shows maximum depression in freezing point?
 (a) 0.5 M Li₂SO₄ (b) 1 M NaCl (c) 0.5 M Al₂(SO₄)₃ (d) 0.5 M BaCl₂
- (v) For a chemical reaction dS = - 0.035 kJ/K and dH = -20 kJ. At what temperature does the reaction turn nonspontaneous?
 (a) 5.14 K (b) 57.14 K (c) 571.4 K (d) 5714.0 K
- (vi) The standard e.m.f. of the following cell is 0.463 V

$$\text{Cu}_{(s)} \left| \text{Cu}_{(aq)}^{++} \right| \left| \text{Ag}_{(aq)}^{++} \right| \text{Ag}_{(s)}$$
 (1M) (1M) . What is the standard potential of Cu electrode?
 (a) 1.137 V (b) 0.337 V (c) 0.463 V (d) -0.463 V
- (vii) Fe₂O₃ is reduced to spongy iron near the top of blast furnace by _____
 (a) H₂ (b) CaO (c) SiO₂ (d) CO

Q. 2. Answer any SIX of the following :

[12]

- (i) Distinguish between crystalline solid and amorphous solid.
 (ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.
 (iii) Write cell reactions in lead storage battery during discharge.
 (iv) Draw structures and write geometry of PCl₃ and PCl₅.
 (v) Prove that ΔH = ΔU + Δ nRT. What is the condition under which ΔU = ΔH?
 (vi) Mention names and formulae of two ores of aluminium.
 (vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.
 (viii) What is pseudo first order reaction? Give one example of it.

Q. 3. Answer any THREE of the following :

[9]

- (i) Calculate the mole fraction and molality of HNO₃ in a solution containing 12.2% HNO₃. (Given - atomic masses : H = 1, N = 14, O = 16)
 (ii) Consider the reaction -



What are the values of the following at the same time ?

- (a) $-\frac{d[\text{I}^{-}]}{dt}$ (b) $-\frac{d[\text{S}_{2}\text{O}_{8}^{2-}]}{dt}$ (c) $-\frac{d[\text{I}_{3}^{-}]}{dt}$

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