

Ques. ► 6. $A = (n-1)d^5ns^1$, $B = (n-1)d^1ns^2$, Here $n = 4$

[D.B.-16]

- What is common-ion effect?
- In which orbital between 3d and 4p electron will enter first?
- Determine the four quantum number of the last electron of the element A.
- Which among the elements A and B will produce colored compound? Explain with reason.

Answer to the question no. 6

[a] The process by which the dissociation rate of the weak electrolyte is reduced due to the presence of a strong electrolyte is known as similar-ion effect.

[b] In which orbital the electron will enter first is determined by Aufbau's principle or $(n + l)$. According to this principle, the orbital of which $(n + l)$ value is low has low energy and electron will enter to it first. But if the $(n + l)$ value of two orbitals are same, then electron will enter the orbital whose n value is lower. For both 3d and 4p orbital the $(n + l)$ value is 5. But for 3d orbital the n value is 3 and for 4p orbital the n value is 4. So, electron will enter 3d orbital first.

[c] The last electron of the given element A enters into $(n-1)d$ orbital.

As, the value of n is 4.

So, the value of four quantum numbers of the last electron are :

(i) Principal quantum number $= n - 1 = 4 - 1 = 3$

(ii) The value of subsidiary quantum number, l for d orbital is 2.

(iii) Magnetic quantum number, $m = 2$.

Ques. ► 7



[D.B.-16]

- Write down the law of mass action.
- Which between H_2SO_3 and HNO_3 is more acidic and why?
- If the MN solution of the stem is saturated then calculate the solubility product of MN at $20^\circ C$.
- If the two solutions in the stem are mixed, will there be any ppt? Explain with reason.

Answer to the question no. 7

[a] At a fixed temperature the rate of any reaction is proportional to the active mass (partial pressure in case of liquid and concentration in case of gas) of its reactants. This is the law of mass action.

[b] Between H_2SO_3 and HNO_3 , HNO_3 is more acidic. Because we know that the acid whose oxidation number of the central atom is more is more acidic.

Here for H_2SO_3 , the oxidation number of the central atom S is

$2x + 3(-2) = 0$
 $2x - 6 = 0$
 $2x = 6$
 $x = 3$

[c] For the reaction $MN \rightleftharpoons M^+ + N^-$
 $K_{sp} = [M^+][N^-]$
 $3.5 \times 10^{-4} = x \cdot x$
 $x = \sqrt{3.5 \times 10^{-4}}$
 $x = 0.0005916$

[d] If the two solutions are mixed, the concentration of MN will be $0.30 M + 0.01 M = 0.31 M$. Since this concentration is greater than the solubility product, a precipitate will form.

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Brinkman**

Hsc Chemistry 1st Paper 24:

Excel HSC Chemistry Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39 95 PAGES 428 pp
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Energy Technology 2014 Cong Wang, Jan de Bakker, Cynthia K. Belt, Animesh Jha, Neale R. Neelameggham, Soobhankar Pati, Leon H. Prentice, Gabriella Tranell, Kyle S. Brinkman, 2013-12-30 This book addresses the issues intricacies and challenges relating to energy and environmental sciences Papers cover various technological aspects of sustainable energy eco systems

and processes that improve energy efficiency and reduce and sequester carbon dioxide CO₂ and other greenhouse emissions. The collection also emphasizes the need for sustainable technologies in extractive metallurgy, materials processing and manufacturing industries with reduced energy consumption and CO₂ emission. Industrial energy efficient technologies include innovative ore beneficiation, smelting technologies, recycling and waste heat recovery. The book also carries contributions from all areas of non nuclear and non traditional energy sources including renewable energy sources such as solar, wind and biomass. Papers from the following symposia are presented in the book: Energy Technologies and Carbon Dioxide Management, High temperature Material Systems for Energy Conversion and Storage, Solar Cell, Silicon

Misconceptions in Chemistry Hans-Dieter Barke, Al Hazari, Sileshi Yitbarek, 2008-11-18. Over the last decades several researchers discovered that children, pupils and even young adults develop their own understanding of how nature really works. These pre concepts concerning combustion, gases or conservation of mass are brought into lectures and teachers have to diagnose and to reflect on them for better instruction. In addition there are school made misconceptions concerning equilibrium, acid base or redox reactions which originate from inappropriate curriculum and instruction materials. The primary goal of this monograph is to help teachers at universities, colleges and schools to diagnose and cure the pre concepts. In case of the school made misconceptions it will help to prevent them from the very beginning through reflective teaching. The volume includes detailed descriptions of class room experiments and structural models to cure and to prevent these misconceptions.

Resources in Education, 1982

Light Metals 2020 Alan Tomsett, 2020-01-28. The Light Metals symposia at the TMS Annual Meeting Exhibition present the most recent developments, discoveries and practices in primary aluminum science and technology. The annual Light Metals volume has become the definitive reference in the field of aluminum production and related light metal technologies. The 2020 collection includes papers from the following symposia: Alumina and Bauxite, Aluminum Alloys, Processing and Characterization, Aluminum Reduction Technology, Cast Shop Technology, Cast Shop Technology, Recycling and Sustainability, Joint Session, Electrode Technology for Aluminum Production.

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