

- Q1** 21. It is necessary to know the formula of a compound in order to calculate its percentage composition.
- AT** 22. If the percent by mass of carbon in methane, CH_4 , is 75%, then 100 grams of methane contains 75 g of percent of hydrogen.
- AT** 23. The formula for methane, CH_4 , is both a molecular and an empirical formula.
- AT** 24. The empirical formula of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, is CH_2O .

C. Problems

Solve the following problems in the space provided. Show your work.

25. How many moles of Fe in 5.1 g of 10^{23} moles of Fe?
26. How many representative particles are in 2.18 g of 10^{23} and MgCl_2 ?
27. What is the gram molecular mass of methane, CH_4 ?
28. Find the mass of each of the following:
- 2.43 g of 10^{23} and H_2O_2
 - 2.43 g and H_2O_2
 - 4.30 g and 1.30 g
 - 4.30 g and 1.30 g

Honors Chemistry Chemical Quantities Answer Key

D Keegan



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