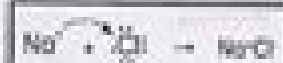


IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:



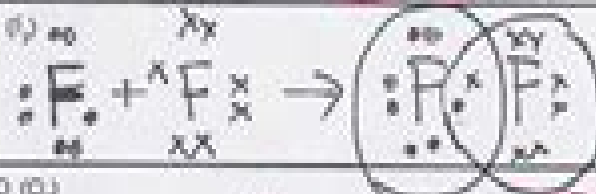
Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO})$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



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1. Understanding the eBook Guided Practice Chaapter 8 Covalent Bonding Answers
 - The Rise of Digital Reading Guided Practice Chaapter 8 Covalent Bonding Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Chaapter 8 Covalent Bonding Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Chaapter 8 Covalent Bonding Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Chaapter 8 Covalent Bonding Answers
 - Personalized Recommendations
 - Guided Practice Chaapter 8 Covalent Bonding Answers User Reviews and Ratings
 - Guided Practice Chaapter 8 Covalent Bonding Answers and Bestseller Lists
5. Accessing Guided Practice Chaapter 8 Covalent Bonding Answers Free and Paid eBooks
 - Guided Practice Chaapter 8 Covalent Bonding Answers Public Domain eBooks
 - Guided Practice Chaapter 8 Covalent Bonding Answers eBook Subscription Services
 - Guided Practice Chaapter 8 Covalent Bonding Answers Budget-Friendly Options

6. Navigating Guided Practice Chapter 8 Covalent Bonding Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Practice Chapter 8 Covalent Bonding Answers Compatibility with Devices
 - Guided Practice Chapter 8 Covalent Bonding Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Practice Chapter 8 Covalent Bonding Answers
 - Highlighting and Note-Taking Guided Practice Chapter 8 Covalent Bonding Answers
 - Interactive Elements Guided Practice Chapter 8 Covalent Bonding Answers
8. Staying Engaged with Guided Practice Chapter 8 Covalent Bonding Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Chapter 8 Covalent Bonding Answers
9. Balancing eBooks and Physical Books Guided Practice Chapter 8 Covalent Bonding Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Chapter 8 Covalent Bonding Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Chapter 8 Covalent Bonding Answers
 - Setting Reading Goals Guided Practice Chapter 8 Covalent Bonding Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Chapter 8 Covalent Bonding Answers
 - Fact-Checking eBook Content of Guided Practice Chapter 8 Covalent Bonding Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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