Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

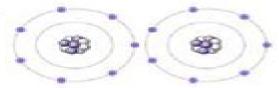
Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- There are eight markers in a full set, but Flora and Frank each only have seven markers.
 Flora is missing the red marker, and Frank is missing the blue marker. What can they do so
 that each has a full set of markers? Frank should give the red marker to Flora, and Flora
 should give the blue marker to Frank
- Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set? Otto should give the black and brown markers to Olivia, and Olivia should give the purple and green markers to Otto

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the Covalent Bonds Gizmo™, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** () to see the electrons orbiting the nucleus of each atom.



- The outermost electrons in each atom are called valence electrons. How many valence electrons does each fluorine atom have? 7 valence electrons
- Click Pause (). Drag an electron from the left atom to the right atom. Click Play.

What happens? They equally share the one atom on the new shell that is formed

Click Pause, drag an electron from the right atom to the left, and then click Play.

What happens now? They continue to spiral around both atoms, but now there are two electrons on the shell

Activity A: Sharing electrons	Get the Gizmo ready	
	Click Reset. Select Hydrogen.	(•)(

Guided Study Covalent Bonds Answer Key

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