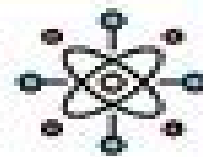


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combinations

1. **IONIC BONDING** occurs between a metal and a nonmetal. The metal atom loses one or more electrons to become a positively charged cation, and the nonmetal atom gains those electrons to become a negatively charged anion. The resulting electrostatic attraction between the oppositely charged ions forms the ionic bond.

2. **COVALENT BONDING** occurs between two nonmetals. The atoms share one or more pairs of electrons, forming a shared electron pair that acts as the glue holding the atoms together.

3. **METALLIC BONDING** occurs between metal atoms. The metal atoms release their valence electrons into a "sea of electrons" that surrounds the positively charged metal cations. The attraction between the cations and the sea of electrons forms the metallic bond.

SCIENCE REAL

A simplified periodic table of elements, color-coded by groups.

IONIC BONDING

When a metal atom (e.g., Sodium, Na) loses an electron, it becomes a positively charged cation (Na⁺). When a nonmetal atom (e.g., Chlorine, Cl) gains an electron, it becomes a negatively charged anion (Cl⁻). The resulting electrostatic attraction between the oppositely charged ions forms the ionic bond.

COVALENT BONDING

When two nonmetal atoms (e.g., Carbon, C) share one or more pairs of electrons, they form a covalent bond. The shared electron pair acts as the glue holding the atoms together.

METALLIC BONDING

When metal atoms (e.g., Copper, Cu) release their valence electrons into a "sea of electrons," the resulting attraction between the cations and the sea of electrons forms the metallic bond.

A diagram showing a sodium cation (Na⁺) and a chloride anion (Cl⁻) with dashed lines representing the electrostatic attraction between them.

GUIDED READING ACTIVITY

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ML Morrison

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