Abbreviations.	1	= alternative and acceptable answers for the same marking point
annotations and	10000	= separates marking points
conventions	NOT	= answers which are not worthy of credit
used in the	()	= words which are not essential to gain credit
	200	<ul> <li>(underlining) key words which <u>must</u> be used to gain credit</li> </ul>
Mark Scheme	ecf.	= error carried forward
	AW	= alternative wording
	ora	= or reverse argument

Question	Expected answers	Marks
1 (a)	8 1, 2 or 3 p orbitals are OK	[2]
(b)	d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓	[3]
	energy $2s$ $\uparrow \downarrow$ $2p$ $\uparrow \downarrow \uparrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\uparrow$ $\downarrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\uparrow$ $\downarrow$ $\downarrow$ $\downarrow$ $\downarrow$ $\uparrow$ $\downarrow$	
	2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓	[2]
(d) (	<ul> <li>Energy change when each atom in 1 mole ✓     of gaseous atoms ✓     loses an electron ✓ (to form 1 mole of gaseous 1+ ions).</li> </ul>	[3]
(	<ul> <li>O<sup>2+</sup>(g) → O<sup>3+</sup>(g) + e<sup>-</sup> ✓ ✓</li> <li>1 mark for correct species; 1 mark for state symbols.</li> <li>No charge required on electron.</li> </ul>	[2]
	Ignore (g) on e	121

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